

6. When one gram each of KMnO_4 reacts with HCl , the volume of chlorine liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres
 (c) 44.8 litres (d) 89.6 litres
7. When MnO_2 reacts with H^+ ions in acid solution Mn^{2+} is liberated. The equivalent weight of KMnO_4 in this reaction is
- (a) 158 (b) 166
 (c) 309 (d) 326
8. 54 g. of hydrogen peroxide is present in 1000 ml. of solution. This solution is called
- (a) 10 vol. solution (b) 20 vol. solution
 (c) 30 vol. solution (d) 50 vol. solution
9. The number of moles of KMnO_4 that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 1/5 (b) 2/5
 (c) 4/5 (d) 1
10. To prepare a solution that is 0.50 M KCl starting with 100 mL of 0.40 M KCl
- (a) add 0.50 g KCl
 (b) add 50 mL of water
 (c) add 0.50 mol KCl
 (d) evaporate 10 mL water
11. In hot acidic solution, Br_2 disproportionates to Br^+ and BrO_2^+
- $$3\text{Br}_2 + 6\text{H}^+ \rightarrow 2\text{Br}^+ + \text{BrO}_2^+ + 3\text{H}_2\text{O}$$
- Atomic equivalent weight of Br_2 is (molecular weight = 160)
- (a) $\frac{160}{3}$ (b) $\frac{160}{2}$
 (c) $\frac{160}{1}$ (d) $\frac{160}{4}$
12. 176 mL of a solution prepared by adding 28.00 mL of 0.02 M MnCl_2 with 70 mL of 0.10 M Na_2SO_4 is
- (a) 0.008 M (b) 0.140 M
 (c) 0.018 M (d) 0.200 M
13. When 80 mL of 0.20 M HCl is mixed with 100 mL of 0.10 M KOH , the resulting solution is the same as a solution of
- (a) 0.10 M HCl and 0.02 M KCl
 (b) 0.08 M HCl
 (c) 0.08 M HCl and 0.01 M KOH
 (d) 0.08 M HCl and 0.01 M HCl
14. 50 mL of H_2O_2 solution contains strength = (x) required 10 mL of KMnO_4 solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.4
 (c) 0.5 (d) 0.6
15. Molarity of 10 M H_2SO_4 (d = 1.8 g/mL) is
- (a) 10 mol kg^{-1} (b) 100 mol kg^{-1}

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Katrin Zwirgmaier



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