

Notes for use in qualitative analysis

Tests for anions

anion	test	test result
carbonate, CO_3^{2-}	add dilute acid, then test for carbon dioxide gas	effervescence, carbon dioxide produced
chloride, Cl^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	white ppt.
bromide, Br^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	cream ppt.
iodide, I^- [in solution]	acidify with dilute nitric acid, then add aqueous silver nitrate	yellow ppt.
nitrate, NO_3^- [in solution]	add aqueous sodium hydroxide, then aluminium foil; warm carefully	ammonia produced
sulfate, SO_4^{2-} [in solution]	acidify with dilute nitric acid, then add aqueous barium nitrate	white ppt.
sulfite, SO_3^{2-}	add a small volume of acidified aqueous potassium manganate(VII)	the acidified aqueous potassium manganate (VII) changes colour from purple to colourless

Tests for aqueous cations

cation	effect of aqueous sodium hydroxide	effect of aqueous ammonia
aluminium, Al^{3+}	white ppt., soluble in excess, giving a colourless solution	white ppt., insoluble in excess
ammonium, NH_4^+	ammonia produced on warming	—
calcium, Ca^{2+}	white ppt., insoluble in excess	no ppt. or very slight white ppt.
chromium(III), Cr^{3+}	green ppt., soluble in excess	green ppt., insoluble in excess
copper(II), Cu^{2+}	light blue ppt., insoluble in excess	light blue ppt., soluble in excess, giving a dark blue solution
iron(II), Fe^{2+}	green ppt., insoluble in excess, ppt. turns brown near surface on standing	green ppt., insoluble in excess, ppt. turns brown near surface on standing
iron(III), Fe^{3+}	red-brown ppt., insoluble in excess	red-brown ppt., insoluble in excess
zinc, Zn^{2+}	white ppt., soluble in excess, giving a colourless solution	white ppt., soluble in excess, giving a colourless solution

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Table of Contents Igcse Chemistry 2mark Scheme Paper 6

1. Understanding the eBook Igcse Chemistry 2mark Scheme Paper 6
 - The Rise of Digital Reading Igcse Chemistry 2mark Scheme Paper 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Chemistry 2mark Scheme Paper 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Chemistry 2mark Scheme Paper 6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Chemistry 2mark Scheme Paper 6

- Personalized Recommendations
 - Igcse Chemistry 2mark Scheme Paper 6 User Reviews and Ratings
 - Igcse Chemistry 2mark Scheme Paper 6 and Bestseller Lists
5. Accessing Igcse Chemistry 2mark Scheme Paper 6 Free and Paid eBooks
 - Igcse Chemistry 2mark Scheme Paper 6 Public Domain eBooks
 - Igcse Chemistry 2mark Scheme Paper 6 eBook Subscription Services
 - Igcse Chemistry 2mark Scheme Paper 6 Budget-Friendly Options
 6. Navigating Igcse Chemistry 2mark Scheme Paper 6 eBook Formats
 - ePub, PDF, MOBI, and More
 - Igcse Chemistry 2mark Scheme Paper 6 Compatibility with Devices
 - Igcse Chemistry 2mark Scheme Paper 6 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Igcse Chemistry 2mark Scheme Paper 6
 - Highlighting and Note-Taking Igcse Chemistry 2mark Scheme Paper 6
 - Interactive Elements Igcse Chemistry 2mark Scheme Paper 6
 8. Staying Engaged with Igcse Chemistry 2mark Scheme Paper 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Chemistry 2mark Scheme Paper 6
 9. Balancing eBooks and Physical Books Igcse Chemistry 2mark Scheme Paper 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Chemistry 2mark Scheme Paper 6
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Igcse Chemistry 2mark Scheme Paper 6
 - Setting Reading Goals Igcse Chemistry 2mark Scheme Paper 6
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Igcse Chemistry 2mark Scheme Paper 6

- Fact-Checking eBook Content of Igcse Chemistry 2mark Scheme Paper 6
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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