

AP CHEMISTRY UNIT 8: EQUILIBRIUM CHANGES: Le Chatelier's Principle

Unit 8 - Introduction to Equilibrium

1. Which of the following is NOT true?

Equilibrium constant for a given reaction does NOT depend on the temperature of the equilibrium system during which it is set.

Equilibrium constant for a reaction is independent of concentration

(a) true

(b) false

(c) both

(d) $\frac{1}{K}$

☒ (e) none

2. Which is NOT a factor affecting the equilibrium constant?

A closed, reacting system subjected to a disturbance. A concentration change shifts the equilibrium without changing temperature. Addition of an inert gas.

A change in temperature of a gas.

A change in volume of a gas mixture.

A change in concentration of a reactant.

(a) A gas

(b) Total Pressure

(c) Total Volume

(d) A gas mixture

☒ (e) All of the above

3. Which is NOT a factor affecting the equilibrium constant?4. Which is NOT a factor affecting the equilibrium constant?

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A reaction involving a mixture of gases in equilibrium with a solid substance in a constant-volume container. Addition of an inert gas shifts the chemical equilibrium of the reaction.

(a) Addition of a reactant to a gas.

(b) Addition of a reactant to a gas.

(c) Adding the product to a gas.

☒ (d) Changing the temperature.

For a reaction, the equilibrium constant is temperature dependent. Which of the following statements is NOT true?

(a) At a fixed T , K is constant.

(b) At a fixed T , K is constant.

(c) At a fixed T , K is constant.

(d) At a fixed T , K is constant.

☒ (e) At a fixed T , K is constant.

(a) true (b) false (c) true (d) true (e) false

5. Which is NOT a factor affecting the equilibrium constant?

Reaction involving a mixture of gases in equilibrium with a solid substance in a constant-volume container. Addition of an inert gas shifts the chemical equilibrium of the reaction.

(a) A gas

(b) Total Pressure

(c) Total Volume

(d) A gas mixture

☒ (e) All of the above

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(c) Total Volume

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