

Setup

English

Cautions and Notes

- Caution:** must be observed to avoid damage to your equipment.
- Note:** contains important information and useful tips on the operation of your printer.

1 Unpacking

- Remove the contents from the box. Remove all tape and packing material from the printer.

Caution:
Be sure not to touch the ribbon cable inside the printer.



- Attach the paper support and plug in the printer.



2 Installing the Ink Cartridges

- Press the **Power** button to turn on the printer.



- Remove the ink cartridges from their protective bags.

Do not touch the ink contacts on the cartridges.



- Install the ink cartridges as shown below.



Note: Install all ink cartridges. The printer will not work unless all cartridges are installed properly.

- Press the **Maintenance** button.

The printer begins ink charging. It takes about **one and a half minutes**. While charging, the printer produces a variety of mechanical sounds. This is normal. When the green light stops flashing, ink charging is complete.



Caution:
Do not turn off the printer during the ink charging process.

Note:
The ink cartridges packaged with your printer are pre-filled during initial setup. In order to produce high quality prints, the print head in your printer will be fully charged with ink. This ink-off process consumes a quantity of ink and subsequent cartridge life will last for their full rated life.

See the other side of this sheet to install the printer software.

Safety Instructions

Read all of the instructions in this section when setting up and using your printer.

When choosing a place for the printer

- Avoid places subject to rapid changes in temperature and humidity. Also, keep the printer away from direct sunlight, direct light, or heat sources.
- Avoid places subject to dust, shocks, or vibrations.
- Leave enough room around the printer to allow for sufficient ventilation.
- Place the printer on a flat, stable surface that extends beyond the base of the printer in all directions. If you place the printer on a desk, leave more than 10 cm between the back of the printer and the desk. The printer will not operate properly if it is tilted or wobbles.
- When storing or transporting the printer, do not tilt it, stand it on its side, or turn it upside down. Otherwise the ink may leak from the cartridges.

When choosing a power source

- Use only the type of power source indicated on the printer's label.
- Use only the power cord that comes with this product. Use of another cord may result in fire or electrical shock.
- The product power cord is for use with this product only. Use with other equipment may result in fire or electrical shock.
- Be sure your AC power cord meets the relevant local safety standard.
- Do not use or damage or bend power cords.
- If you use an extension cord with the printer, make sure that the total ampere rating of the device plugged into the extension cord does not exceed the cord's ampere rating. Also, make sure that the extension cord is rated for the voltage of the wall outlet that you are using.
- If you plan to use the printer in Germany, the building installation must be protected by a fuse to drop circuit breaker to provide adequate short-circuit protection and over-current protection for the printer.

When handling ink cartridges

- Keep the cartridges out of the reach of children. Do not allow children to tamper with or otherwise handle the cartridges.
- Do not open the ink cartridge packages until just before you install them.
- Do not shake the ink cartridges. This can cause leakage.
- Be careful when you handle used ink cartridges, as there may be some ink remaining around the ink supply port. If the ink gets on your skin, wash the area thoroughly with soap and water. If it gets into your eyes, flush them immediately with water. If it comes into contact with your clothes, wash them immediately. Do not touch the yellow tape seal from the cartridges before installing them. Otherwise, the cartridges will become unusable.
- Do not remove the blue portion of the tape seal from the top of the cartridges.
- Do not remove the tape seal from the bottom of the cartridges.
- The IC chip on the ink cartridge retains a variety of cartridge-related information, such as the amount of ink remaining, so that a cartridge may be removed and reinstalled freely. However, each time the cartridge is inserted, some ink is consumed because the printer automatically performs a reliability check.
- Do not touch the green IC chip on the side of the cartridge.

When using the printer

- Do not put your hand inside the printer or touch the ink cartridges during printing.
- Do not block or cover the vent and openings in the printer.
- Do not attempt to service the printer yourself.
- Unplug the printer and refer servicing to qualified service personnel under the following conditions: The power cord/plug is damaged; the plug does not fit properly in the printer; the printer has been dropped or the casing is damaged; the printer does not operate normally or exhibits a distinct change in performance.
- Do not insert objects through the slots in the printer.
- Take care not to spill liquid on the printer.
- Store the ink cartridges in a cool, dry place. Removing the cartridges can dehydrate the print head and may prevent the printer from printing.

Epson Stylus C45 Manual

Epson (Firm)



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web the purpose of the copper iron stoichiometry lab was to determine which of the two possible iron sulfate compounds formed when reacting aqueous copper ii sulfate with solid iron metal the techniques used in this lab were weighing by difference quantitative transfer and vacuum filtration

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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75 and higher the coordination of trimethylamine by copper ii

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g

