

## Chemistry Honors Worksheet: Formulas of a Compound

1.
  - a.) moles The subscripts in a chemical formula represent the ratio of ( moles / mass ) in a compound
  - b.) CO Which compound has the greatest percentage of carbon? (CO, CO<sub>2</sub>, PbCO<sub>3</sub> )
  - c.) CH<sub>4</sub> What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
  - d.) 83 g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
  - e.) Same The % composition of an empirical formula is ( more / less / the same ) than/as its molecular formula.
  - f.) one How many moles of V are indicated in one mole of the compound V(C<sub>2</sub>O<sub>4</sub>)<sub>2</sub>?
  - g.)  $1 \times 10^{24}$  atoms A beaker of water contains  $2 \times 10^{24}$  atoms of H, how many atoms of O are also in the beaker?
  - h.)  $6.02 \times 10^{23}$  1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
  - i.) 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

2. The pesticide DDT has been banned in the United States since 1962. This compound has a formula C<sub>14</sub>H<sub>9</sub>Cl<sub>5</sub>. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl 709.5 g/mol

- a.) molecular Is this compound ionic or molecular?
- b.)  $5.2 \times 10^{15}$  How many molecules of DDT are in a  $6.1 \times 10^{-4}$  g sample?
- c.)  $1.53 \times 10^{17}$  atoms H How many atoms of H are in 2.00 mg of DDT?
- d.)  $7.513$  Cl How many grams of Cl are contained in 10.0 grams of DDT?
- e.)  $7.84 \times 10^{21}$  atoms C If a sample  $5.04 \times 10^{21}$  atoms of H, how many atoms of C would there also be?
- f.)  $2.97 \times 10^{20}$  atoms C How many atoms of C are in a 25.0 mg of DDT?

3. A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

|                                      |         |            |               |
|--------------------------------------|---------|------------|---------------|
| Mass of crucible:                    | 24.22 g | > 1.72g Ta | 0.0106 mol Ta |
| Mass of crucible and tantalum:       | 26.14 g |            |               |
| Mass of crucible and tantalum oxide: | 26.96 g | > 0.42g O  | 0.02625 mol O |

- a.) Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- b.) What is the empirical formula of this compound? TaO<sub>2.5</sub> - Ta<sub>2</sub>O<sub>5</sub>
- c.) Is this compound ionic or covalent? ionic
- d.) Name this compound: Tantalum(V) oxide \* Need Roman Numerals!!!

4. Determine the empirical formula of the following compounds:

- a.) 63.0g Rb, 5.90g O Rb<sub>2</sub>O
- b.) 4.95mg Th, 1.37mg S ThS<sub>2</sub>
- c.) 2.13g Na, 2.32g As, 1.98g O Na<sub>3</sub>AsO<sub>4</sub>
- d.) 32.8% Cr, 67.2% Cl CrCl<sub>3</sub>
- e.) 58.0% Rb, 9.50% N, 32.5% O RbNO<sub>3</sub>

# Empirical Molecular Formula Answers

**JS Bruner**



## **Empirical Molecular Formula Answers:**

## **Empirical Molecular Formula Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Empirical Molecular Formula Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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