

Experiment 7 Report Sheet

Empirical Formulas

Date _____ Lab Sec _____ Name _____

Indicate whether the following data are for Part B, Part C, or Part D. _____

Desk No. _____

Identify the original sample _____

	Trial 1	Trial 2
1. Mass of crucible and lid (g)	25.636	40.347
2. Mass of crucible, lid, and sample (g)	25.786	40.542
3. Mass of sample (g) <u>0.150</u>	0.150	0.195
4. Instructor's approval of flame and apparatus	—	—
5. Mass of crucible, lid, and product		
1st mass measurement (g)	25.858	40.654
2nd mass measurement (g)	25.856	40.652
3rd mass measurement (g)	25.859	40.658
6. Final mass of crucible, lid, and product (g)	25.860	40.655
7. Mass of product (g)	0.224g	0.113g
8. Part B. Combination Reaction of Magnesium and Oxygen		
a. Mass of O (g)	0.014g	
b. Mass ratio of Mg to O	2:1	
c. Moles of Mg (mol)	0.00617 mol	
d. Moles of O (mol)	0.00463 mol	
e. Mole ratio of Mg to O	1:1	
f. Consensus empirical formula of magnesium oxide	<u>1 MgO</u>	
g. Percent by mass (%): <u>67.0</u> % Mg; <u>33.0</u> % O		

9. Part C. Decomposition Reaction of a Pure Compound

a. Mass of product = mass CaO (g)	0.224	
b. Mass of CO ₂ (g)	0.1263	

e) note

Empirical Formula Of A Compound Lab Report

Jo A. Beran



Empirical Formula Of A Compound Lab Report:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 *Laboratory Manual for Principles of General Chemistry* 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section.

Basic Laboratory Principles in General Chemistry Fitzgerald B. Bramwell, 1990 **Laboratory Manual** Jo A. Beran, 1990 **Bibliography of Scientific and Industrial Reports**, 1947 **Merck's Report** Theodore Weicker, 1910 **The Merck Report**, 1910 *Introduction to Chemistry, Laboratory Manual* T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic, provocative, and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text: developing chemical thinking, a chemical vision, learning problem-solving methods, and utilizing group work and discussion activities. These themes involve and engage the students in their own learning processes; they are challenged to be active. The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics. The reorganization presents many core concepts in the first five chapters, preparing students for later chapters. In addition, the author has added vignettes throughout the chapters referring to health, technology, the environment, and society, as well as to specific tools of direct use to students.

Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998 **Experimental Chemistry** Wilbert Hutton, 1977 *Scientific and Technical Aerospace Reports*, 1988 *Nationalizing Science* Alan J. Rocke, 2000-11-08 After looking at the early careers of Wurtz's two mentors, Liebig and Jean Baptiste Dumas, Rocke describes Wurtz's life and career in the politically complex period leading up to 1853. He then discusses the turning point in Wurtz's intellectual life: his conversion to the reformed chemistry of Laurent Gerhardt and Williamson and his efforts to persuade his colleagues of the

advantages of the new system In 1869 Adolphe Wurtz 1817 1884 called chemistry a French science In fact however Wurtz was the most internationalist of French chemists Born in Strasbourg and educated partly in the laboratory of the great Justus Liebig he spent his career in Paris where he devoted himself to introducing German ideas into French scientific circles His life therefore provides an excellent vehicle for considering the divergent trajectories of French and German chemistry and by extension French and German science during this crucial period After looking at the early careers of Wurtz s two mentors Liebig and Jean Baptiste Dumas Rocke describes Wurtz s life and career in the politically complex period leading up to 1853 He then discusses the turning point in Wurtz s intellectual life his conversion to the reformed chemistry of Laurent Gerhardt and Williamson and his efforts social and political as well as scientific to persuade his colleagues of the advantages of the new system He looks at political patronage or the lack thereof and at the insufficient material support from the French government during the middle decades of the century From there Rocke goes on to examine the rivalry between Wurtz and Marcellin Berthelot the debate over atoms versus equivalents and the reasons for Wurtz s failure to win acceptance for his ideas The story offers insights into the changing status of science in this period and helps to explain the eventual course of both French and German chemistry

Operational Organic Chemistry John W. Lehman, 1999 For sophomore level organic lab courses This text lab manual helps students master the fundamental laboratory operations of organic chemistry and develop critical thinking skills through scientific problem solving

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