

Chemistry Honors Worksheet: Formulas of a Compound

1.
 - a.) moles The subscripts in a chemical formula represent the ratio of (moles / mass) in a compound
 - b.) CO Which compound has the greatest percentage of carbon? (CO, CO₂, PbCO₃)
 - c.) CH₄ What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
 - d.) 83 g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
 - e.) Same The % composition of an empirical formula is (more / less / the same) than/as its molecular formula.
 - f.) one How many moles of V are indicated in one mole of the compound V(C₂O₄)₂?
 - g.) 1×10^{24} atoms A beaker of water contains 2×10^{24} atoms of H, how many atoms of O are also in the beaker?
 - h.) 6.02×10^{23} 1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
 - i.) 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

2. The pesticide DDT has been banned in the United States since 1962. This compound has a formula C₁₄H₉Cl₅. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl

- a.) molecular Is this compound ionic or molecular?
- b.) 5.2×10^{15} How many molecules of DDT are in a 6.1×10^{-4} g sample?
- c.) 1.53×10^{17} atoms H How many atoms of H are in 2.00 mg of DDT?
- d.) 7.51 g Cl How many grams of Cl are contained in 10.0 grams of DDT?
- e.) 7.84×10^{21} atoms C If a sample 5.04×10^{21} atoms of H, how many atoms of C would there also be?
- f.) 2.97×10^{20} atoms C How many atoms of C are in a 25.0 mg of DDT?

3. A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

Mass of crucible:	24.22 g		
Mass of crucible and tantalum:	26.14 g	> 1.92 g Ta	0.0106 mol Ta
Mass of crucible and tantalum oxide:	26.56 g	> 0.42 g O	0.02625 mol O

- a.) Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- b.) What is the empirical formula of this compound? TaO_{2.5} - Ta₂O₅
- c.) Is this compound ionic or covalent? ionic
- d.) Name this compound: Tantalum(V) oxide * Need Roman Numerals!!!

4. Determine the empirical formula of the following compounds:

- a.) 63.0g Rb, 5.90g O Rb₂O
- b.) 4.95mg Th, 1.37mg S ThS₂
- c.) 2.13g Na, 2.32g As, 1.98g O Na₃AsO₄
- d.) 32.8% Cr, 67.2% Cl CrCl₃
- e.) 58.0% Rb, 9.50% N, 32.5% O RbNO₃

Empirical Molecular Formula Answers

Chandan Sengupta



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An Introduction to Chemistry Michael Mosher,Paul Kelter,2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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