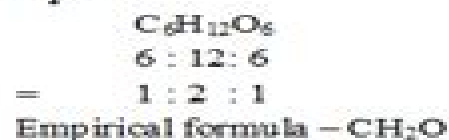
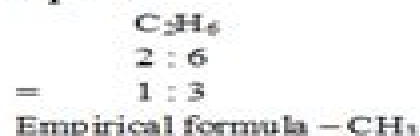


**Empirical formula:** It shows the simplest whole number ratio between the atoms of the elements in the compound.

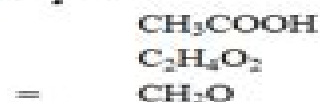
**Example 1:**



**Example 2:**



**Example 3:**



**Practice:**

1.  $\text{C}_4\text{H}_8$
2.  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
3.  $\text{N}_2\text{O}_4$

[www.icsechemistry16.blogspot.com](http://www.icsechemistry16.blogspot.com)

**Calculate the empirical formula of the compounds having the following percentage composition:**

**Example 1:**

C-40%  
H-6.7%  
O-53.3%

**Sol:**

Element	%	% ÷ At. Wt.	Simplest ratio	simplest whole number ratio
Carbon	40	$40 \div 12 = 3.33$	$\frac{3.33}{3.33} = 1$	<b>1</b> C
Hydrogen	6.7	$6.7 \div 1$	$\frac{6.7}{3.33} = 2$	<b>2</b> H
Oxygen	53.3	$53.3 \div 16$	$\frac{3.33}{3.33} = 1$	<b>1</b> O
Empirical Formula		<b>= CH<sub>2</sub>O</b>		

# Empirical Formula Practice

**Lei Huang**

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