

How to Calculate an Empirical Formula

A 283.9 g sample of unknown composition was carefully analyzed and found to consist of 49.8 g of sodium, 113.6 g of chromium, and 121.4 g of oxygen. Determine the empirical formula of the unknown compound.

① Convert mass to moles

② Divide By Lowest Amount of moles

$$\frac{49.8 \text{ g Na}}{22.99 \text{ g Na}} \times \frac{1 \text{ mol Na}}{1} = \frac{2.166 \text{ mol Na}}{2.166} = \boxed{1} \times 2$$

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$$\frac{113.6 \text{ g Cr}}{52.00 \text{ g Cr}} \times \frac{1 \text{ mol Cr}}{1} = \frac{2.185 \text{ mol Cr}}{2.166} = \boxed{1} \times 2$$

$$\frac{121.4 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{7.588 \text{ mol O}}{2.166} = \boxed{3.5} \times 2 =$$

Empirical Formula Determination

Michael Hoelscher



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Table of Contents Empirical Formula Determination

1. Understanding the eBook Empirical Formula Determination
 - The Rise of Digital Reading Empirical Formula Determination
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Determination
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Determination
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Determination
 - Personalized Recommendations
 - Empirical Formula Determination User Reviews and Ratings
 - Empirical Formula Determination and Bestseller Lists
5. Accessing Empirical Formula Determination Free and Paid eBooks

- Empirical Formula Determination Public Domain eBooks
- Empirical Formula Determination eBook Subscription Services
- Empirical Formula Determination Budget-Friendly Options
- 6. Navigating Empirical Formula Determination eBook Formats
 - ePub, PDF, MOBI, and More
 - Empirical Formula Determination Compatibility with Devices
 - Empirical Formula Determination Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Determination
 - Highlighting and Note-Taking Empirical Formula Determination
 - Interactive Elements Empirical Formula Determination
- 8. Staying Engaged with Empirical Formula Determination
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Formula Determination
- 9. Balancing eBooks and Physical Books Empirical Formula Determination
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Determination
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Determination
 - Setting Reading Goals Empirical Formula Determination
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Determination
 - Fact-Checking eBook Content of Empirical Formula Determination
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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