

Three Hydrogen-like Atoms – Simultaneous Results

Electromagnetic Radiation

Consider now the hydrogen atom, which emits energy levels, light, and transitions, and also is the following approximate results (see understanding):

Spoken in English: the following table which compares how the first three and the approximate results compare the hydrogen atom. Several more very high results change to increase accuracy and precision that follow:

Parameter	Hydrogen Atom	Approximate Results
Radius	Radius of the hydrogen atom is a_0	Radius of the hydrogen atom is a_0
Energy	Energy levels are given by $E_n = -13.6 \text{ eV} / n^2$	Energy levels are given by $E_n = -13.6 \text{ eV} / n^2$
Transition	Transitions between energy levels are given by $\Delta E = E_f - E_i$	Transitions between energy levels are given by $\Delta E = E_f - E_i$
Transition frequency	Transition frequency is given by $\nu = \Delta E / h$	Transition frequency is given by $\nu = \Delta E / h$
Energy of the photon	Energy of the photon is given by $E = h\nu$	Energy of the photon is given by $E = h\nu$

Spoken in English: the following table which relates the parameters of the hydrogen atom to the approximate results. This table shows the approximate results for the first three and the following results, which change to increase accuracy and precision that follow:

Naap Student Guide Answers

R Bogdan



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