



## Lab Data



Verify your calculation. In your set-up, do your units cancel?

Mass of magnesium (g)	.015
Moles of magnesium (mol)	.00062
Temperature of water (°C)	23.0
Temperature of water (K)	296.15
Vapor pressure of water (mmHg)	21.1
Barometric pressure (mmHg)	778.0

## Observations

I observe that there are bubbles that are visible.

Volume of hydrogen gas collected (mL)	15.00
Volume of hydrogen gas collected (L)	0.01500
Pressure of hydrogen gas (mmHg)	756.9
Pressure of hydrogen gas (atm)	0.9959
Moles of hydrogen gas produced (mol)	.00690
Ideal gas constant (L · atm · mol <sup>-1</sup> · K <sup>-1</sup> )	0.080

# Ideal Gas Constant Lab 38 Answers

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**Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

*Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      **PGT Physics Exam PDF-Physics**

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