7

Chemical Equilibrium

1. For the reversible reaction.

 $N_2(g) + 3H_2(g) \longrightarrow 2NH_2(g) + heat$ the equilibrium shifts in forward direction

- (a) by increasing the concentration of NH₂(g)
- (b) by decreasing the pressure
- (c) by decreasing the concentrations of N₂(g) and H₂(g)
- (d) by increasing pressure and decreasing temperature
- Using the Gibbs energy change, ΔG°=+ 63.3 kJ for the following reaction,

 $Ag_2CO_3(s) \longrightarrow 2Ag^*(aq) + CO_3^{2-}(aq)$

the $K_{\rm sp}$ of Ag₂ CO₃(s) in water at 25°C is $(R = 8.314 \, {\rm JK^{-1} mol^{-1}})$ [2014]

- (a) 3.2×10^{-26}
- (b) 8.0×10^{-12}
- (c) 2.9×10^{-3}
- (d) 7.9×10^{-2}
- KMnO₄ can be prepared from K₂MnO₄ as per reaction.

3MnO₄^{2−} + 2H₂O ----

The reaction can go to completion by removing OH⁻ ions by adding [2013]

- (a) HCI
- (b) KOH
- (c) CO₂
- (d) SO₂

The value of ΔH for the reaction,

$$X_2(g) + 4Y_2(g) - 2XY_4(g)$$

is less than zero. Formation of $XY_4(g)$ will be favoured at [2011]

- (a) low pressure and low temperature
- (b) high temperature and low pressure
- (c) high pressure and low temperature
- (d) high temperature and high pressure
- 5. For the reaction,

 $N_2(g) + O_2(g) \longrightarrow 2NO(g)$, the equilibrium constant is K_1 . The equilibrium constant is K_2 for the reaction.

 $2NO(g) + O_2(g) \longrightarrow 2NO_2(g)$. What is K for the reaction, $NO_2(g) \longrightarrow \frac{1}{2}N_2(g) + O_2(g)$?

(a) 1/(4 K₁K₂)

(b) [1/K₁K₂]^{5/2} [2011]

(c) 1/(K,K2)

- (d) 1/(2K,K₂)
- 6. In which of the following equilibrium K_c and K_p are not equal? [2010]
 - (a) $2NO(g) \longrightarrow N_2(g) + O_2(g)$
 - (b) $SO_2(g) + NO_2(g) \longrightarrow SO_3(g) + NO(g)$
 - $(c) H_2(g) + I_2(g) \longrightarrow 2HI(g)$
 - (d) 2C(s) + O₂(g) → 2 CO₂(g)
- The dissociation constants for acetic acid and HCN at 25°C are 1.5 × 10⁻⁵ and 4.5 × 10⁻¹⁰, respectively. The equilibrium constant for the equilibrium,

CN⁻ + CH₃COOH → HCN + CH₃COOwould be [2009]

(a) 3.0×10^5

(b) 3.0×10^{-5}

(c) 3.0×10^{-4}

(d) 3.0×10^4

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RC Schank

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Table of Contents Equilibrium Questions With Answers

- 1. Understanding the eBook Equilibrium Questions With Answers
 - The Rise of Digital Reading Equilibrium Questions With Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Equilibrium Questions With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Questions With Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Equilibrium Questions With Answers
 - Personalized Recommendations
 - Equilibrium Questions With Answers User Reviews and Ratings
 - Equilibrium Questions With Answers and Bestseller Lists
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 - Highlighting and Note-Taking Equilibrium Questions With Answers
 - Interactive Elements Equilibrium Questions With Answers
- 8. Staying Engaged with Equilibrium Questions With Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Equilibrium Questions With Answers
- 9. Balancing eBooks and Physical Books Equilibrium Questions With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Questions With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Equilibrium Questions With Answers
 - Setting Reading Goals Equilibrium Questions With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Equilibrium Questions With Answers
 - Fact-Checking eBook Content of Equilibrium Questions With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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