

 **K_{sp} and ICE Problems Explained**

Key: K_{sp} = solubility product
 ICE = Initial, Change, Equilibrium

1. Write the K_{sp} expression for silver chloride using the following solubility product equation:



Initial	K_{sp}	Change	Equilibrium
AgCl	1.8×10^{-10}	$+x$	x
Ag ⁺	x	$+x$	x
Cl ⁻	x	$+x$	x

2. Write the K_{sp} expression for silver chloride using the following solubility product equation and solve for the solubility of silver chloride.



Initial	K_{sp}	Change	Equilibrium
AgCl	1.8×10^{-10}	$+x$	x
Ag ⁺	x	$+x$	x
Cl ⁻	x	$+x$	x

3. Find the K_{sp} for silver chloride if the solubility of silver chloride is 1.3×10^{-5} mol/L. Write the K_{sp} expression for silver chloride.



If the solubility of silver chloride is 1.3×10^{-5} mol/L, then the concentration of silver ions is 1.3×10^{-5} mol/L and the concentration of chloride ions is 1.3×10^{-5} mol/L.

Initial	K_{sp}	Change	Equilibrium	K_{sp}	Equilibrium
AgCl	1.8×10^{-10}	$+x$	x	x	x
Ag ⁺	x	$+x$	x	x	x
Cl ⁻	x	$+x$	x	x	x

Equilibrium Ice Box Answers

Martin William Lisse



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