

## **SPECTROPHOTOMETRIC DETERMINATION OF AN EQUILIBRIUM CONSTANT**

### **Lab report**

Name:

ID #:

TA name:

Lab Section:

Lab Date:

1. Looking at Table 5 (data sheet), how is  $K_c$  related to the initial concentrations (  $[Fe^{3+}]_0$  and  $[SCN^-]_0$ )? Is it what you expected? (Explain briefly) (5 marks)

Looking at Table 5 in the results, it is somewhat evident that  $K_c$  is actually independent of the initial concentrations of the reactants since they are relatively consistent among the five test tubes. Furthermore, the  $K_c$  values are much more relative as opposed to the significant changes in  $K_c$  due to temperature variation which holds true because the equilibrium constant is affected by temperature.

However, looking at the data for any trends, it appears that  $K_c$  drops as the concentration of  $[SCN^-]$  increases, but this should not be true. So, accounting for relative human error and limited machine capabilities, the value only fluctuated minimally from the average which was 204 (M<sup>-1</sup>). Thus the results are not exactly as I had expected, but ignoring the small deviations, the results are based on Le Chatelier's Principle which states that when a chemical system in equilibrium is disturbed there will be a shift in the system to reduce the effect of that disturbance. So, by increasing  $[SCN^-]$  from tube 1 to 5, there should be no change of  $K_c$  since equilibrium will shift to the right of the reaction to increase the concentration of  $[FeSCN^{2+}]$  and balance out.

$$K_c = \frac{[FeSCN^{2+}]_{eq}}{[Fe^{3+}] * [SCN^-]}$$

2. To obtain  $[FeSCN_{2-}]_{eq}$  in tube 5, you made the assumption that 100% of the ions  $SCN^-$  had reacted. Now that you know the value for  $K_c$  (average in table 5), build an ICE diagram to calculate the true percentage of  $SCN^-$  reacted in tube 5. Comment on the validity of the assumption. (Make sure you include the ICE diagram in your answer) (10 marks)

ICE Table

	$[Fe^{3+}]$	$+ [SCN^-]$	$\rightleftharpoons [FeSCN_{2-}]$
Initial	0.180 M	0.000199 M	0 M
Change	-x	-x	+x
Equilibrium	$(0.180 - x)$ M	$(0.000199 - x)$ M	x M

# Equilibrium Constant Lab Report Discussion

**Elaine Seymour**



## **Equilibrium Constant Lab Report Discussion:**

The book delves into Equilibrium Constant Lab Report Discussion. Equilibrium Constant Lab Report Discussion is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Equilibrium Constant Lab Report Discussion, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Equilibrium Constant Lab Report Discussion
- Chapter 2: Essential Elements of Equilibrium Constant Lab Report Discussion
- Chapter 3: Equilibrium Constant Lab Report Discussion in Everyday Life
- Chapter 4: Equilibrium Constant Lab Report Discussion in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Equilibrium Constant Lab Report Discussion. The first chapter will explore what Equilibrium Constant Lab Report Discussion is, why Equilibrium Constant Lab Report Discussion is vital, and how to effectively learn about Equilibrium Constant Lab Report Discussion.

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6. In chapter 5, the author will draw a conclusion about Equilibrium Constant Lab Report Discussion. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Equilibrium Constant Lab Report Discussion.

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