

SPECTROPHOTOMETRIC DETERMINATION OF AN EQUILIBRIUM CONSTANT

Lab report

Name:

ID #:

TA name:

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Lab Date:

1. Looking at Table 5 (data sheet), how is K_c related to the initial concentrations ($[Fe^{3+}]_0$ and $[SCN^-]_0$)? Is it what you expected? (Explain briefly) (5 marks)

Looking at Table 5 in the results, it is somewhat evident that K_c is actually independent of the initial concentrations of the reactants since they are relatively consistent among the five test tubes. Furthermore, the K_c values are much more relative as opposed to the significant changes in K_c due to temperature variation which holds true because the equilibrium constant is affected by temperature.

However, looking at the data for any trends, it appears that K_c drops as the concentration of $[SCN^-]$ increases, but this should not be true. So, accounting for relative human error and limited machine capabilities, the value only fluctuated minimally from the average which was 204 (M⁻¹). Thus the results are not exactly as I had expected, but ignoring the small deviations, the results are based on Le Chatelier's Principle which states that when a chemical system in equilibrium is disturbed there will be a shift in the system to reduce the effect of that disturbance. So, by increasing $[SCN^-]$ from tube 1 to 5, there should be no change of K_c since equilibrium will shift to the right of the reaction to increase the concentration of $[FeSCN^{2+}]$ and balance out.

$$K_c = \frac{[FeSCN^{2+}]_{eq}}{[Fe^{3+}] * [SCN^-]}$$

2. To obtain $[FeSCN_{2-}]_{eq}$ in tube 5, you made the assumption that 100% of the ions SCN^- had reacted. Now that you know the value for K_c (average in table 5), build an ICE diagram to calculate the true percentage of SCN^- reacted in tube 5. Comment on the validity of the assumption. (Make sure you include the ICE diagram in your answer) (10 marks)

ICE Table

	$[Fe^{3+}]$	$+ [SCN^-]$	$\rightleftharpoons [FeSCN_{2-}]$
Initial	0.180 M	0.000199 M	0 M
Change	-x	-x	+x
Equilibrium	$(0.180 - x)$ M	$(0.000199 - x)$ M	x M

Equilibrium Constant Lab Report

**Ralph Petrucci, William
Harwood, Geoffrey Herring**



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Table of Contents Equilibrium Constant Lab Report

1. Understanding the eBook Equilibrium Constant Lab Report
 - The Rise of Digital Reading Equilibrium Constant Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Constant Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Constant Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Constant Lab Report
 - Personalized Recommendations

- Equilibrium Constant Lab Report User Reviews and Ratings
- Equilibrium Constant Lab Report and Bestseller Lists
- 5. Accessing Equilibrium Constant Lab Report Free and Paid eBooks
 - Equilibrium Constant Lab Report Public Domain eBooks
 - Equilibrium Constant Lab Report eBook Subscription Services
 - Equilibrium Constant Lab Report Budget-Friendly Options
- 6. Navigating Equilibrium Constant Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Constant Lab Report Compatibility with Devices
 - Equilibrium Constant Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Constant Lab Report
 - Highlighting and Note-Taking Equilibrium Constant Lab Report
 - Interactive Elements Equilibrium Constant Lab Report
- 8. Staying Engaged with Equilibrium Constant Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Constant Lab Report
- 9. Balancing eBooks and Physical Books Equilibrium Constant Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Constant Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Equilibrium Constant Lab Report
 - Setting Reading Goals Equilibrium Constant Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Equilibrium Constant Lab Report
 - Fact-Checking eBook Content of Equilibrium Constant Lab Report

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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