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## Ch 11 Worksheet 5 - Thermochemical Stoichiometry

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1. Given the equation:  $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l})$   $\Delta \text{H} = -571.7 \text{ kJ}$
- How much energy is produced if 100.0 grams of  $\text{H}_2$  are converted to water?
  - How many grams of  $\text{H}_2$  would be required to produce 5000.0 kJ?
2. Given the equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$   $\Delta \text{H}_{\text{rxn}} = -92.22 \text{ kJ}$
- Is the formation of ammonia endothermic or exothermic?
  - What is the enthalpy change associated with the formation of 1.0 mol of  $\text{NH}_3$ ?
  - Is the *decomposition* of ammonia endothermic or exothermic?
  - What is the enthalpy change associated with the decomposition of 4.00 mol  $\text{NH}_3$ ?
  - How much energy will be released when 12.0 g  $\text{H}_2(\text{g})$  react with excess  $\text{N}_2(\text{g})$  to form  $\text{NH}_3(\text{g})$ ?
3. Given the equation:  $2\text{Ca}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{CaO}(\text{s})$   $\Delta \text{H} = -635.09 \text{ kJ}$
- How much energy is released during the production of 5.0 g of  $\text{CaO}$ ?
  - How much energy must be added to  $\text{CaO}$  to produce 64.0 g of  $\text{O}_2(\text{g})$ ?
4. Given the equation:  $4\text{Fe}(\text{s}) + 3 \text{O}_2(\text{g}) \rightarrow 2\text{Fe}_2\text{O}_3(\text{s})$   $\Delta \text{H} = -824.2 \text{ kJ}$
- Assuming that the iron is converted to iron (III) oxide as in the above equation, how much heat can be provided by a hot pack that contains 50.0 g of iron?
  - If the hot pack from part a. gives off heat at the rate of 42.0 kJ /hour, how long will it stay warm?

Answers:

1a. 14,150 kJ produced

1b. 35.3g

2a.

2b. -46kJ

2c.

2d. 184 kJ

2e. 183 kJ

3a. 28 kJ released

3b. 1270 kJ

4a. 185 kJ provided

4b. 4.40 hours

# Energy Ws 5 Thermochemical Equations And Stoichiometry

**JR Anderson**



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