

Let the relative atomic mass of element Z be  $z$ .

Element	Z	Bromine, Br
Mass (g)	2.07	$3.67 - 2.07 = 1.60$
Number of moles of atoms	$\frac{2.07}{z}$	$\frac{1.60}{80} = 0.02$

Based on the empirical formula  $ZBr_2$ , the ratio of atoms of Z : Br is 1 : 2.

Therefore,  $\frac{2.07}{z} : 0.02 = 1 : 2$

$$\frac{2.07}{z} \div 0.02 = 1 \div 2$$

$$\frac{2.07}{0.02z} = \frac{1}{2}$$

$$z = \frac{2 \times 2.07}{0.02}$$

$$= 207$$

The relative atomic mass of element Z is 207.

# Empirical Formula For Copper Oxide

**Marco Cascella**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

**Empirical Formula For Copper Oxide:**

## The Enigmatic Realm of **Empirical Formula For Copper Oxide**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Empirical Formula For Copper Oxide** a literary masterpiece penned by a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those who partake in its reading experience.

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