

Empirical and Molecular Formula Practice Solutions

1. A compound is found to contain 63.52 % iron and 36.48% sulfur. Find its empirical formula.

$$63.52 \text{ g Fe} \times \frac{1 \text{ mol}}{55.85 \text{ g}} = 1.137 \text{ mol Fe}$$

$$36.48 \text{ g S} \times \frac{1 \text{ mol}}{32.06 \text{ g}} = 1.138 \text{ mol S}$$

$$\frac{1.137 \text{ mol Fe}}{1.137} : \frac{1.138 \text{ mol S}}{1.137} = 1 \text{ mol Fe: } 1.001 \text{ mol S}$$

empirical formula = FeS

2. A compound is found to contain 26.56% potassium, 35.41% chromium, and the remainder oxygen. Find its empirical formula.

$$26.56 \text{ g K} \times \frac{1 \text{ mol}}{39.10 \text{ g}} = 0.6793 \text{ mol K}$$

$$36.48 \text{ g Cr} \times \frac{1 \text{ mol}}{52.00 \text{ g}} = 0.7015 \text{ mol Cr}$$

$$36.96 \text{ g O} \times \frac{1 \text{ mol}}{16.00 \text{ g}} = 2.310 \text{ mol O}$$

$$\frac{0.6793 \text{ mol K}}{0.6793} : \frac{0.7015 \text{ mol Cr}}{0.6793} : \frac{2.310 \text{ mol O}}{0.6793} = 1 \text{ mol K: } 1.033 \text{ mol Cr: } 3.401 \text{ mol O}$$

$$\rightarrow 2 \text{ mol K: } 2.066 \text{ mol Cr: } 6.802 \text{ mol O}$$

empirical formula = $\text{K}_2\text{Cr}_2\text{O}_7$

3. A 60.00 g sample of tetraethyllead, a gasoline additive, is found to contain 38.43 g lead, 17.83 g carbon, and 3.74 g hydrogen. Find its empirical formula.

$$38.43 \text{ g Pb} \times \frac{1 \text{ mol}}{207.20 \text{ g}} = 0.1855 \text{ mol Pb}$$

$$17.83 \text{ g C} \times \frac{1 \text{ mol}}{12.01 \text{ g}} = 1.485 \text{ mol C}$$

$$3.74 \text{ g H} \times \frac{1 \text{ mol}}{1.01 \text{ g}} = 3.78 \text{ mol H}$$

$$\frac{0.1855 \text{ mol Pb}}{0.1855} : \frac{1.485 \text{ mol C}}{0.1855} : \frac{3.78 \text{ mol H}}{0.1855} = 1 \text{ mol Pb: } 8.005 \text{ mol C: } 20.4 \text{ mol H}$$

empirical formula = $\text{PbC}_8\text{H}_{20}$

Empirical And Molecular Formula Calculation Practice

David Kirk



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