# Experiment 7 Report Sheet

Em	A	The second second	a House of
			具 医甲基

Land whether the Car		Empirical Formulas
Infaction whether the following data are for Part B, p	in c	Denk No.
	or Part D.	Identify the original sample
<ol> <li>Mass of crucible and lid (g)</li> </ol>	Trial I	Trial 2
2. Mass of crucible, lid, and sample (g)	-25_636	_ 40.347
3. Mass of sample (g) D. 150	-25.786	_ 40.542
4. Instructor's approval of flame and apparatus	_0.150	0.195
5. Mass of crucible, lid, and product		
1st mass measurement (g)	0-0-0	
2nd mass measurement (g)	25.858	_ 40.654
	25.856	40.652
3rd mass measurement (g)	25.859	40.656
6. Final mass of crucible, lid, and product (g)	25.860	40.655
7. Mass of product (g)	0.2249	0.1139
S. Part B. Combination Reaction of Magnesium ar	ad Oxygen	
a. Mass of O (g)	0.014a	
h. Mass ratio of Mg to O	2:1	
c. Moles of Mg (mol)	0.00617 mot	
d. Moles of O (mol)	0.00463 mol	
e. Mole ratio of Mg to O	1:1	
Consensus empirical formula of magnesium oxid	6	1/Mg0
Percent by mass (%): 67.0 % Mg; 3		
art C. Decomposition Reaction of a Pure Compo	ound	
Mass of product = mass CaO (g)	0.224	The second secon
Mass of CO <sub>2</sub> (g)	0.1263	to the second second

# **Empirical Formula Of A Compound Lab Report**

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# **Empirical Formula Of A Compound Lab Report:**

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best and get the college credits you deserve a Advanced Placement Program and AP are registered trademarks of the College Board which was not involved in the production of and does not endorse this product Modern Experimental Chemistry George W. Jr. Latimer, 2012-12-02 Modern Experimental Chemistry provides techniques of qualitative analysis that reinforce experiments on ionic equilibriums This book includes the determination of water in hydrated salts identification of an organic compound after determining its molecular weight and nonaqueous titration of a salt of a weak acid The calculation of chemical stoichiometry calculation of thermodynamic properties by determining the change in equilibrium with temperature and chromium chemistry are also covered This compilation contains enough experiments for classes which have six hours of laboratory two 3 hour meetings per week to last two semesters This publication is intended for chemistry students as an introductory manual to chemistry laboratory **Introduction to Chemistry, Laboratory Manual** T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities. These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics. The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment **Government Reports Announcements & Index**, 1976 and society as well as to specific tools of direct use to students

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political patronage or the lack thereof and at the insufficient material support from the French government during the middle decades of the century From there Rocke goes on to examine the rivalry between Wurtz and Marcellin Berthelot the debate over atoms versus equivalents and the reasons for Wurtz s failure to win acceptance for his ideas The story offers insights into the changing status of science in this period and helps to explain the eventual course of both French and German chemistry

Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi Arthur D.

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