Empirical and Molecular Formulas Worksheet



Objectives:

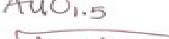
· be able to calculate empirical and molecular formulas

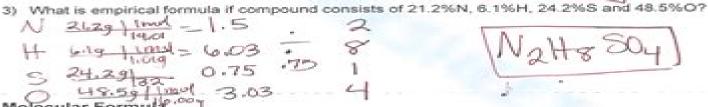
Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?



2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?





4) Empirical formula of a substance is CH₂O. Molar mass is 180. What is the molecular formula?



5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

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JS Bruner

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