

## EMPIRICAL AND MOLECULAR FORMULAE

**EMPIRICAL FORMULA** - This represents the simplest whole number ratio of atoms within a molecule

**MOLECULAR FORMULA** - This represents the exact numbers of each atom in a molecule

Empirical formulae can be calculated from either % mass composition data or mass composition data. Molecular formula can then be calculated if the molecular mass of the compound is known.

### EXAMPLE 1 - from % composition

A substance containing carbon and hydrogen and oxygen only contains 73.47% carbon and 10.20% hydrogen by mass, its relative molecular mass is 98. Calculate its molecular formula.

Step 1 calculate empirical formula

ATOMS	CARBON	HYDROGEN	OXYGEN
PERCENTAGE MASS	73.47	10.20	$(100 - (73.47 + 10.20)) = 16.33$
ATOMIC MASS	12	1	16
MASS/ATOMIC MASS	6.1225	10.20	1.020
RATIO(divide by smallest)	6	10	1

Empirical formula =  $C_6H_{10}O$

Step 2 use  $M_r$  to find molecular formula

$$M_r = \text{mass } (C_6H_{10}O) \times n$$

$$98 = (98) \times n$$

$$n = 1$$

molecular formula =

$C_6H_{10}O$

### EXAMPLE 2 - from mass data

A substance contains 12.28g nitrogen, 3.51g hydrogen, 28.07g sulphur and 56.10g oxygen, its  $M_r$  is 228. Calculate its molecular formula.

# Empirical And Molecular Answers And Questions

**SB Merriam**



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