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O García



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Diesel Models 670 770 870 970 1070 **Technical Service Manual** , *TE 20 Workshop and Service Manual* Harry
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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{products} - \Delta H_{reactants}$ is positive if heat is absorbed that is if $\Delta H_{products}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in

solution eventually allowing an individual to prove or disprove Hess's Law. Introduction: Thermodynamics is the study of energy exchanged between the system and the surroundings.

beyondlabz 5 Hess's Law.docx 3 12 Hess's Law in this - Mar 10 2023

Web 3 12 Hess's Law: In this experiment, you will measure the amount of heat released in these three related exothermic reactions:

1. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{aq} + \text{OH}^- \text{aq}$
2. $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{Cl}^- \text{aq} + \text{H}_2\text{O (l)}$
3. $\text{NaOH (s)} + \text{H}^+ \text{aq} \rightarrow \text{Cl}^- \text{aq} + \text{H}_2\text{O (l)}$