

Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

- 31 What is the number of electrons in an Al^{3+} ion? Low 5
Al +13p
- (1) 10 (3) 3
(2) 13 (4) 16

- 32 The valence electron of which atom in the ground state has the greatest amount of energy?

- ps. 6 (1) cesium ps. 5 (3) rubidium
ps. 2 (2) lithium ps. 3 (4) sodium

- 33 The numbers of protons and neutrons in each of four different atoms are shown in the table below.

Way #23
Protons and Neutrons in Four Different Atoms

Atom	Number of Protons	Number of Neutrons
A	8	8
D	9	9
E	9	10
G	10	10

Which two atoms represent isotopes of the same element?

- (1) A and D (3) E and D
(2) A and G (4) E and G

- 34 Which elements have the most similar chemical properties?

- Way #1
38
(1) boron and carbon
(2) oxygen and sulfur
(3) aluminum and bromine
(4) argon and silicon

- 35 Which element reacts with oxygen to form ionic bonds?

- Way #44
(1) calcium (3) chlorine
(2) hydrogen (4) nitrogen

- 36 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of chlorine.

Naturally Occurring Isotopes of Chlorine

Isotopes	Atomic Mass of the isotope (u)	Natural Abundance (%)
^{35}Cl	34.97	75.76
^{37}Cl	36.97	24.24

Way #24
Which numerical setup can be used to calculate the atomic mass of the element chlorine?

- (1) $(34.97 \text{ u})(75.76) + (36.97 \text{ u})(24.24)$
(2) $(34.97 \text{ u})(0.2424) + (36.97 \text{ u})(0.7576)$
(3) $(34.97 \text{ u})(0.7576) + (36.97 \text{ u})(0.2424)$
(4) $(34.97 \text{ u})(24.24) + (36.97 \text{ u})(75.76)$

- Way #42
37 Which general trends in first ionization energy and electronegativity values are demonstrated by Group 15 elements as they are considered in order from top to bottom?

- (1) The first ionization energy decreases and the electronegativity decreases.
(2) The first ionization energy increases and the electronegativity increases.
(3) The first ionization energy decreases and the electronegativity increases.
(4) The first ionization energy increases and the electronegativity decreases.

- 38 An aluminum sample has a mass of 80.01 g and a density of 2.70 g/cm³. According to the data, to what number of significant figures should the calculated volume of the aluminum sample be expressed?

- (1) 1 (3) 3
(2) 2 (4) 4

$$D = \frac{M}{V} \quad ; \quad V = \frac{M}{D}$$

$$\frac{80.01 \text{ (4)}}{2.70 \text{ (3)}} \text{ (3)}$$

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Table of Contents Jan Regents 2015 Answers Chemistry

1. Understanding the eBook Jan Regents 2015 Answers Chemistry
 - The Rise of Digital Reading Jan Regents 2015 Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Jan Regents 2015 Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jan Regents 2015 Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Jan Regents 2015 Answers Chemistry
 - Personalized Recommendations
 - Jan Regents 2015 Answers Chemistry User Reviews and Ratings
 - Jan Regents 2015 Answers Chemistry and Bestseller Lists
5. Accessing Jan Regents 2015 Answers Chemistry Free and Paid eBooks
 - Jan Regents 2015 Answers Chemistry Public Domain eBooks
 - Jan Regents 2015 Answers Chemistry eBook Subscription Services
 - Jan Regents 2015 Answers Chemistry Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Jan Regents 2015 Answers Chemistry Compatibility with Devices
 - Jan Regents 2015 Answers Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jan Regents 2015 Answers Chemistry
 - Highlighting and Note-Taking Jan Regents 2015 Answers Chemistry
 - Interactive Elements Jan Regents 2015 Answers Chemistry
8. Staying Engaged with Jan Regents 2015 Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jan Regents 2015 Answers Chemistry
9. Balancing eBooks and Physical Books Jan Regents 2015 Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jan Regents 2015 Answers Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Jan Regents 2015 Answers Chemistry
 - Setting Reading Goals Jan Regents 2015 Answers Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Jan Regents 2015 Answers Chemistry
 - Fact-Checking eBook Content of Jan Regents 2015 Answers Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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