

The Hydrogen Atom – Student Challenge

Background Material

Carefully read the background-information Energy Levels, Light, and Transitions, and answer the following questions as clearly (and completely) as you can.

Operation 1: Compare the following table which compares how the Bohr Model and the Quantum-mechanical approach the Hydrogen atom. In each “describing” box describe things as accurately as you can remember they are.

Concept	Bohr Model	Quantum Model
Electrons	positively charged particles in flat orbits	positively charged particles in flat orbits
Electrons	orbiting electrons	negatively charged clouds
Orbitals	Electrons do not move in any flat orbitally paths	Electrons form constant orbital waves
Transitions of electrons to the next orbital	moving a distance to the next orbital	Quantum jumps
Energy of the lowest level	$E_1 = -13.6 \text{ eV}$	$E_1 = -13.6 \text{ eV}$

Operation 2: Compare the following table which compares the properties of two different elements by relating their approximate relationships. The data now is completed for you. The red element has a higher frequency, smaller frequency, smaller energy, and the same wavelength. Which space is missing please?

Hydrogen Atom Student Guide Solutions Naap

Karen S. Brewer



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