



Part III: Balance the following chemical equations

20. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{AlCl}_3$
21. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{FeO} \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{O}_3 + \underline{\hspace{1cm}} \text{Fe}$
22. $\underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{NaCl}$
23. $\underline{\hspace{1cm}} \text{P}_4\text{O}_{10} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{H}_3\text{PO}_4$
24. $\underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{CaSO}_4$
25. $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
26. $\underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
27. $\underline{\hspace{1cm}} \text{N}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{HNO}_3$
28. $\underline{\hspace{1cm}} \text{H}_3\text{AsO}_4 \rightarrow \underline{\hspace{1cm}} \text{As}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
29. $\underline{\hspace{1cm}} \text{Pb} + \underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{C}_2\text{H}_5\text{Cl} \rightarrow \underline{\hspace{1cm}} \text{Pb}(\text{C}_2\text{H}_5)_4 + \underline{\hspace{1cm}} \text{NaCl}$

Part IV: In each of the following problems you are given some potential reaction. Determine what type of reaction would occur, predict the products, write the complete chemical equation and balance it!

30. Aluminum powder burns in the presence of oxygen to form ...
31. Magnesium metal reacts with nitric acid to form ...
32. Aqueous barium chloride reacts with sodium hydroxide to form ...
33. Butane burns in the presence of oxygen to form ...
34. Hydrogen gas reacts with chlorine gas to form ...
35. Aluminum powder reacts with iron(III) oxide to form ...

Formulas And Chemical Reactions Study Guide

Daniel F McAuley



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Chemistry Equations And Answers (Speedy Study Guides) Speedy Publishing, 2015-04-24 Chemistry is a difficult subject to fully comprehend with its equations and scientific laws Trying to digest an entire book in one semester is a tough job but with the help of study guides like these you can absorb information in chemistry much more effectively This guide covers chemical equations including examples potential problems and solutions **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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