

NAME _____

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GRAPHING PERIODIC TRENDS

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

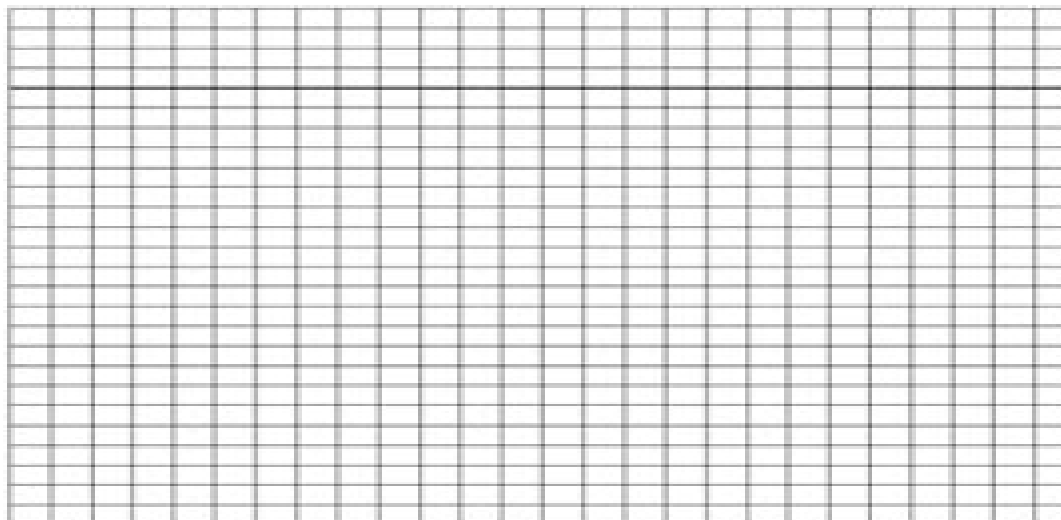
First ionization energy-

Electronegativity-

Graph 1 – Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to **trace** over the element's period (horizontal row on the periodic table). For example: use GREEN to **trace** for all of the elements in row 1, then use YELLOW to **trace** for all of the elements in row 2, then use ORANGE to **trace** for all the elements in row 3, then use BLUE to **trace** for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)
H	1	37
He	2	50
Li	3	123
Be	4	89
B	5	80
C	6	77
N	7	70
O	8	66
F	9	64
Ne	10	67
Na	11	157
Mg	12	136
Al	13	125
Si	14	117
P	15	110
S	16	104
Cl	17	99
Ar	18	98
K	19	203
Ca	20	174



1. Describe the trend in atomic radius as the atomic number increases across a period.
2. Describe the trend in atomic radius as the atomic number increases down a group.

Graphing Periodic Trends Worksheet Answers

DJ Losen



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Table of Contents Graphing Periodic Trends Worksheet Answers

1. Understanding the eBook Graphing Periodic Trends Worksheet Answers
 - The Rise of Digital Reading Graphing Periodic Trends Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Graphing Periodic Trends Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Graphing Periodic Trends Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Graphing Periodic Trends Worksheet Answers
 - Personalized Recommendations
 - Graphing Periodic Trends Worksheet Answers User Reviews and Ratings
 - Graphing Periodic Trends Worksheet Answers and Bestseller Lists
5. Accessing Graphing Periodic Trends Worksheet Answers Free and Paid eBooks

- Graphing Periodic Trends Worksheet Answers Public Domain eBooks
 - Graphing Periodic Trends Worksheet Answers eBook Subscription Services
 - Graphing Periodic Trends Worksheet Answers Budget-Friendly Options
6. Navigating Graphing Periodic Trends Worksheet Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Graphing Periodic Trends Worksheet Answers Compatibility with Devices
 - Graphing Periodic Trends Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Graphing Periodic Trends Worksheet Answers
 - Highlighting and Note-Taking Graphing Periodic Trends Worksheet Answers
 - Interactive Elements Graphing Periodic Trends Worksheet Answers
8. Staying Engaged with Graphing Periodic Trends Worksheet Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Graphing Periodic Trends Worksheet Answers
9. Balancing eBooks and Physical Books Graphing Periodic Trends Worksheet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Graphing Periodic Trends Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Graphing Periodic Trends Worksheet Answers
- Setting Reading Goals Graphing Periodic Trends Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Graphing Periodic Trends Worksheet Answers
- Fact-Checking eBook Content of Graphing Periodic Trends Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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