

## CHAPTER # 2

## GRAHAM'S LAW OF DIFFUSION

QUESTION # 24, pg. 72, STBB

Helium gas takes 5 seconds to effuse from a hole of 10 dm<sup>3</sup> container, how long it would take for 10 dm<sup>3</sup> of oxygen gas to effuse from the same container at the same temperature and pressure?

### DATA:

$$t_{He} = 5 \text{ s}$$

$$M_{He} = 4 \text{ amu}$$

$$M_{O_2} = 32 \text{ amu}$$

$$t_{O_2} = ?$$

$$\frac{t_A}{t_B} = \sqrt{\frac{M_A}{M_B}}$$

Direct Relation

Finding  $t_{O_2}$  :

Hint: (The substance which is required should taken as numerator)

Using Graham's Law

$$\frac{t_{O_2}}{t_{He}} = \sqrt{\frac{M_{O_2}}{M_{He}}}$$

$$\frac{t_{O_2}}{5} = \sqrt{\frac{32}{4}}$$

$$\frac{t_{O_2}}{t_{He}} = \sqrt{8}$$

$$\frac{t_{O_2}}{5} = 2.828$$

$$t_{O_2} = 2.828 \times 5$$

$$t_{O_2} = 14.14 \text{ s}$$

Ends...

# Grahams Law Of Diffusion Answer Key

**K Payea**



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