NUMERICALS (CHEM-I)

CHAPTER # 2 GRAHAM'S LAW OF DIFFUSION



Dr. ZAIN UL ABDEEN

QUESTION # 24, pg. 72, STBB

Hint: (The substance which is required

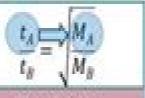
Helium gas takes 5 seconds to effuse from a hole of 10 dm3 container, how long it would take for 10 dm3 of oxygen gas to effuse from the same container at the same temperature and pressure?

DATA:

$$t_{He} = 5 \text{ s}$$

$$M_{He} = 4 \text{ amu}$$

$$M_{O_2} = 32 \text{ amu}$$



Direct Relation

Finding t_{0_2} :

Using Graham's Law

$$\frac{t_{O_2}}{t_{He}} = \sqrt{\frac{M_{O_2}}{M_{He}}}$$

$$\frac{t_{O_2}}{5} = \sqrt{\frac{32}{4}}$$

$$\frac{t_{O_2}}{t_{He}} = \sqrt{8}$$

$$\frac{t_{0_2}}{5} = 2.828$$

$$t_{O_2} = 2.828 \times 5$$

$$t_{\theta_2} = 14.14 \, s$$

Ends...

Grahams Law Of Diffusion Answer Key

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Grahams Law Of Diffusion Answer Key:

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