

Gay-Lussac's Law

Ex. (3) What would be the new temperature of a rigid container of a gas at 101.3 kPa and 293 K if the gas pressure increased to 115 kPa?

$$\frac{\cancel{P_1} P_1}{T_1} = \frac{\cancel{P_2} P_2}{T_2}$$

$$\frac{(101.3 \text{ kPa})}{(293 \text{ K})} = \frac{(115 \text{ kPa})}{(X)}$$

$$X = 333 \text{ K}$$

Gay Lussac S Worksheet Answers

F Rizvi



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