

## Calculating Specific Heat Worksheet

Name: \_\_\_\_\_

Date: \_\_\_\_\_



$Q = mc\Delta T$ , where  $Q$  = heat energy,  $m$  = mass, and  $\Delta T$  = change in temp. Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ . Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$Q = (m)(c)(\Delta T)$$

$$1086.75 = (15.75)(c)(175 - 25)$$

$$1086.75 = 2362.5c$$

$$\frac{1086.75}{2362.5} = \frac{2362.5c}{2362.5}$$

$$c = 0.4672^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = (m)(c)(\Delta T)$$

$$Q = (10)(0.90)(55 - 22)$$

$$Q = 297\text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$Q = (m)(c)(\Delta T)$$

$$67500 = (1500)(c)(57 - 32)$$

$$67500 = 37500c$$

$$\frac{67500}{37500} = \frac{37500c}{37500}$$

$$c = 1.8\text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mc\Delta T$$

$$Q = (100)(4.18)(37 - 4)$$

$$Q = 13827\text{ J}$$

# Myfoundationslab Answer Key

**Daniel F McAuley**



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