

# REPORT SHEET

  

## Heat of Neutralization

EXPERIMENT  
28

### A. Heat Capacity of Calorimeter

1. Temp. of calorimeter and water before mixing
2. Temp. of warm water
3. Maximum temp. determined from your curve
4. Heat lost by warm water (temp decrease  $\times$  50.0 g  $\times$  4.184 J/K-g) =

$$\begin{array}{r} 26 \\ 39 \\ 41 \end{array} \text{ } ^\circ\text{C}$$
  

$$57.822 \text{ J}$$

5. Heat gained by cooler water (temp. increase  $\times$  50.0 g  $\times$  4.184 J/K-g) =

$$58.861 \text{ J}$$

6. Heat gained by the calorimeter [(4) - (5)] =

$$-1047 \text{ J}$$

7. Heat capacity of calorimeter:

$$\frac{\text{heat gained by the calorimeter}}{\text{temperature increase}}$$

$$-3.72 \text{ J/K}$$

### B. Heat of Neutralization of HCl-NaOH

1. Temp. of calorimeter and NaOH  
Temp. of HCl
2.  $\Delta T$  determined from your curve after adding HCl to the NaOH

$$\begin{array}{r} 25 \\ 25 \\ 8 \end{array} \text{ } ^\circ\text{C}$$

3. Heat gained by solution (temperature increase  $\times$  100 g  $\times$  4.184 J/K-g) =

$$117570.4 \text{ J}$$

4. Heat gained by calorimeter (temperature increase  $\times$  heat capacity of calorimeter) =

$$1045.32 \text{ J}$$

5. Total joules released by reaction [(3) + (4)] =

$$116,525 \text{ J}$$

# Heat Of Neutralization Report Sheet Answers

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Nordulf**



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