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## Chapter 5: Thermal Energy

### Lesson 1: Thermal Energy, Temperature, and Heat

**Thermal energy** – The sum of kinetic and potential energy of the particles that make up a material

- Thermal energy causes a phase change.
- Thermal energy describes the energy of the particles that make up a solid, liquid, or gas.

#### **Thermal Energy – Kinetic**

- Every solid, liquid, gas contains trillions of tiny particles that are in motion.
- Because these particles are in motion they all have kinetic energy.
- The faster the particles move, the more kinetic energy they have.
  - Solids – particles vibrate in place
  - Liquids – particles move a little more freely
  - Gases – particles move most freely and quickly

#### **Thermal Energy – Potential**

- Particles that make up solids, liquids, and gases also have potential energy which is the distance between particles.
- The greater the distance between particles the greater the potential energy of the particles.
  - Solids – Particles are held close together by attractive forces
  - Liquids – Particles are slightly farther apart
  - Gases – Particles are much more spread out than in those in a solid

**Heat** – The movement of thermal energy from a warmer object to a cooler object

Why do your hands get warm when you hold a cup of coffee? The thermal energy from the warm cup is moving to and heating up your cool hands.

# Heat Energy Study Guide Answer Key

**G Orfield**



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