

Heat Problems I

Standards:

- 3.1.10 B Describe the concepts of models as a way to predict and understand science and technology.
- 3.4.10 B Analyze energy sources and transfers of heat.

Specific Heat Capacities of some Common Substances

Substance	Cp (J/g°C)	Substance	Cp (J/g°C)
Water (l)	4.184	Iron (s)	0.45
Water (s)	2.03	Mercury (l)	0.14
Water (g)	2.0	Carbon (s)	0.71
Aluminum (s)	0.89	Silver (s)	0.24
		Gold (s)	0.13

$$1\text{Cal} = 4.184\text{Joules}$$

$$1\text{kJ} = 1000\text{J}$$

Problems:

- Convert the following number of calories into joules. (Use D.A. and sig figs.)
 - 100.0 cal = _____ J

b. 1.00×10^3 cal = _____ J

- Convert the following numbers of joules (J) into kilojoules (kJ). (Use D.A. and sig figs.)

a. 243,000 J = _____ kJ

b. 0.251 J = _____ kJ

Heat And Temperature Answer Sheet

Liying Dong

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Heat And Temperature Answer Sheet:

Asian Physics Olympiad (1st - 8th) Yongling Zheng, 2010 This book compiles all of the test problems and solutions from the 1st through the 8th Asian Physics Olympiad Test questions of every paper consist of two parts a theory section and an experiment section before which minutes of teams and results of each competition are introduced It is a rather desirable reference book for both students and teachers of international competition training as well as middle school student contestants

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