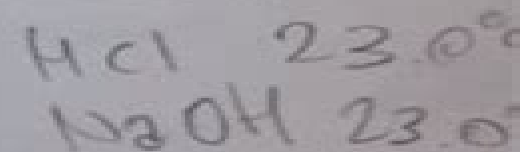


	Reaction 1	Reaction 2	Reaction 3
1. Initial temperature	22.0°C	23.0°C	23.0°C
2. Final temperature	24.0°C	21.0°C	26.0°C
3. The change in temperature, ΔT	+2.0°C	-2.0°C	3.0°C
4. Heat, Q, absorbed by the solution = mass of soln. $\times \Delta T \times 1 \text{ cal/g } ^\circ\text{C}$ Note: Mass = sum of masses of solids and liquids used.	$151.6 \times 2 = 303.2 \text{ cal}$	$160 \text{ ml} \times 4.184 \text{ J/g}^\circ\text{C} \times 2.0^\circ\text{C} = 1300 \text{ cal}$	$166.15 - 9.27 = 156.88 \text{ g} \times 3.0^\circ\text{C} \times 1 \text{ cal/g}^\circ\text{C} = 470.64 \text{ cal}$
5. The amount of heat absorbed by the cup (neglect this quantity)	neglect	neglect	neglect
6. TOTAL amount of heat, Q, absorbed	303.2 cal	640.0 cal	470.64 cal
7. The number of moles of NaOH used in each reaction For solid: $\text{gNaOH} \times 1/\text{MM NaOH}$ For solution: $\text{L NaOH} \times \text{Molarity}$	0.04 mol	0.04 mol	0.04 mol
8. The heat of reaction per mole of NaOH, ΔH . Include sign. (Line 6/Line 7)	7,580 cal	16,000 cal	11,766 cal

(FOR EACH)

Questions

1. Write the net ionic equations for reactions 1, 2, and 3.



2. In reaction 1, ΔH_1 represents the heat of solution of NaOH(s). Look at the net ionic equations for reactions 2 and 3 and make a statement concerning the significance of ΔH_2 , that is, what has happened to produce the energy change. Also explain the meaning of ΔH_3 .

3. How does ΔH_2 compare with the sum of $\Delta H_1 + \Delta H_3$. Explain.

4. Calculate the percent difference between ΔH_2 and the SUM of $\Delta H_1 + \Delta H_3$, using ΔH_2 as the reference.

$$\frac{\text{Difference}}{\Delta H_2} \times 100\% =$$

_____ % difference

Heat Reaction Lab Answers

Yas Asghari

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