

# Orbitals and Bonding Test #12 SCH4U Name: \_\_\_\_\_

     K        T        C        A  
15     10     8     6

## Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (15K)

1. What is the electron configuration of potassium?

- A)  $1s^2 2s^2 2p^2 3s^2 3p^2 4s^2$                       C)  $1s^2 2s^2 3s^2 3p^6 3d^2$   
B)  $1s^2 2s^2 2p^{10} 3s^2 3p^3$                       D)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

2. Which electron configuration represents a violation of the Pauli exclusion principle?



3. Which one of the following sets of quantum numbers can correctly represent a 3p orbital?

- |            |            |            |            |            |
|------------|------------|------------|------------|------------|
| A) $n = 3$ | B) $n = 1$ | C) $n = 3$ | D) $n = 3$ | E) $n = 3$ |
| $l = 1$    | $l = 3$    | $l = 2$    | $l = 1$    | $l = 0$    |
| $m_l = 2$  | $m_l = 3$  | $m_l = 1$  | $m_l = -1$ | $m_l = 1$  |

4. How many electrons in an atom can have  $n = 4$ ,  $l = 2$ ?

- A) 6                      B) 14                      C) 10                      D) 2                      E) None of these

# General Chemistry Ii Chapter Covalent Bonding Orbitals

## Practice Test

**Arshad Iqbal**



## **General Chemistry II Chapter Covalent Bonding Orbitals Practice Test:**

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