

TOPICS:

- ❖ Intermolecular Forces
- ❖ Colligative Properties

Intermolecular Forces

→ The attraction between molecules

Intermolecular forces are always due to coulombic attraction, which is the attraction between molecules due to electric charge



Ion -Dipole

- Ion and polar molecule interact (Polar; unequal pull between electrons, often has lone pairs on central atoms.)

Hydrogen Bonding

- Looking for bonds between hydrogen and fluorine, oxygen, and nitrogen

Dipole-Dipole

- Polar molecules

London Dispersion/Van der Waals

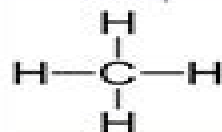
- Mainly nonpolar molecules

Arrange in increasing order: CH_4 , H_2S , CCl_4 , NH_3

Step 1 - draw lewis structures

Step 2 - determine intermolecular forces

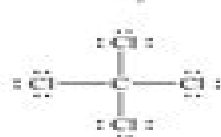
london dispersion



dipole-dipole



london dispersion



hydrogen bonding



Increasing Order: CH_4 , CCl_4 , H_2S , NH_3

*The larger the molar mass, the stronger the intermolecular force. Use this as a tiebreaker

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1. Understanding the eBook Gen Chem 151 Final Exam Review Guide
 - The Rise of Digital Reading Gen Chem 151 Final Exam Review Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Gen Chem 151 Final Exam Review Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gen Chem 151 Final Exam Review Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gen Chem 151 Final Exam Review Guide
 - Personalized Recommendations
 - Gen Chem 151 Final Exam Review Guide User Reviews and Ratings
 - Gen Chem 151 Final Exam Review Guide and Bestseller Lists
5. Accessing Gen Chem 151 Final Exam Review Guide Free and Paid eBooks
 - Gen Chem 151 Final Exam Review Guide Public Domain eBooks
 - Gen Chem 151 Final Exam Review Guide eBook Subscription Services
 - Gen Chem 151 Final Exam Review Guide Budget-Friendly Options

6. Navigating Gen Chem 151 Final Exam Review Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Gen Chem 151 Final Exam Review Guide Compatibility with Devices
 - Gen Chem 151 Final Exam Review Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gen Chem 151 Final Exam Review Guide
 - Highlighting and Note-Taking Gen Chem 151 Final Exam Review Guide
 - Interactive Elements Gen Chem 151 Final Exam Review Guide
8. Staying Engaged with Gen Chem 151 Final Exam Review Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gen Chem 151 Final Exam Review Guide
9. Balancing eBooks and Physical Books Gen Chem 151 Final Exam Review Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gen Chem 151 Final Exam Review Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gen Chem 151 Final Exam Review Guide
 - Setting Reading Goals Gen Chem 151 Final Exam Review Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gen Chem 151 Final Exam Review Guide
 - Fact-Checking eBook Content of Gen Chem 151 Final Exam Review Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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