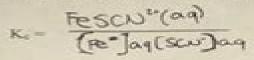
#### Report Sheet - Experiment 13-Chemical Equilibrium

- 1. Write the chemical equation for the reversible reaction studied in this experiment:

  Fe \*(aq)+ SCN \* > Fe (SCN) \*(aq)
- 2. Write the equilibrium expression for the reaction studied in this experiment:



3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added Fe(NO <sub>1</sub> ) <sub>1</sub>	Reddish color	dorker
2	Added KSCN	Reddish color	darker
3	Added NH <sub>4</sub> CI	Reddish color	Lighter
4	Added AgNO	Reddish color	Lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	Fesch to a mix	Shifted Aykt	principle but react on from left to right accust was when more reactant add
3	[SCN] was	Shifted aight	as more reactant was added, the concentration of the product goes up
4	Fezz concentration was	Shifetell Just+	Fe COL WILL Form which leads to a lower Fe to concentration

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**Brendan G. Carr** 

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