

*Lab (25 pts):  
Empirical Formula  
of Magnesium Oxide*

Name \_\_\_\_\_  
Partner \_\_\_\_\_  
Period \_\_\_\_\_ Date \_\_\_\_\_

**Purpose:** To experimentally determine the empirical formula for magnesium oxide,  $\text{Mg}_x\text{O}_y$ .

**Introduction:**

By now you should know that a **molecular formula** is a group of elemental symbols and subscript numbers, which represent the composition of a molecule. The molecular formula shows you how many of each atom can be found in a certain molecule. Some examples of molecular formulas, that you are most likely familiar with, are shown in Table A.

Table A - Some Molecular Formulas and the Information they Convey	
Formula	Composition
$\text{H}_2\text{O}$	Total of 3 atoms in each molecule - 1 oxygen and 2 hydrogen
$\text{O}_2$	Total of 2 atoms in each molecule - 2 oxygen atoms
$\text{CO}_2$	Total of 3 atoms in each molecule - 1 carbon and 2 oxygen
$\text{C}_6\text{H}_{12}\text{O}_6$	Total of 24 atoms in each molecule - 6 carbon, 12 hydrogen, and 6 oxygen.

While the molecular formula shows the molecule as it actually exists, the **empirical formula** for a compound shows the simplest whole number ratio for the elements in the compounds, like reducing a fraction. To find the empirical formula for a compound, find the largest whole number that can be divided into each of the subscripts evenly (greatest common factor). If no number greater than 1 will go into any of the subscripts evenly, then the empirical formula and the molecular formula for the compound are the same. Table B shows examples of some empirical and molecular formulas.

Table B - Comparing Empirical and Molecular Formulas			
Compound	Molecular Formula	GCF	Empirical Formula
Water	$\text{H}_2\text{O}$	1	$\text{H}_2\text{O}$
Hydrogen Peroxide	$\text{H}_2\text{O}_2$	2	HO
Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$	6	$\text{CH}_2\text{O}$
Methane	$\text{CH}_4$	1	$\text{CH}_4$
Ethane	$\text{C}_2\text{H}_6$	2	$\text{CH}_3$
Octane	$\text{C}_8\text{H}_{18}$	2	$\text{C}_4\text{H}_9$

In this lab, you will perform the following reaction:



where you will determine the subscripts,  $x$  and  $y$ . The mass of the solid (magnesium) will increase as the oxygen gas combines with magnesium metal to form magnesium oxide. You will measure the mass of Mg before the reaction and the mass of  $\text{Mg}_x\text{O}_y$  after the reaction to determine the mass of O that reacted from the air, and the mass of Mg and O in magnesium oxide. Then you will calculate magnesium oxide's empirical formula.

**Procedure:** Goggles are **ABSOLUTELY** required!! Don't touch a **HOT** crucible--it will burn you!

- 1) Obtain a stainless steel crucible and lid, and record the mass of the crucible and lid together (watch units & sig figs). Slight discoloration of the crucible is normal and will not affect your results. We are using special balanced for this lab, which measure mass to 0.001 g. *Record all masses to 0.001 g (3 decimal pts).* Use gloves or a crucible tong to handle the crucible. When carrying the crucible hold it firmly with crucible tongs, and hold a piece of metal gauze beneath the crucible in case it slips.



# Lab Determining An Empirical Formula Mgo Answers

**RoseMarie Gallagher, R. Gallagher, Paul  
Ingram**

## Lab Determining An Empirical Formula Mgo Answers:

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web for example take the reaction of lead ii nitrate with hydrochloric acid to form lead ii chloride and nitric acid shown below  
 $\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3$

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