

**DATA TABLE.** Please record your conductivity values in the table provided.

Solution	Concentration	Conductivity ( $\mu\text{S}/\text{cm}$ )
$\text{S} \times \text{A} - \text{CaCl}_2$		10110 $\mu\text{S}/\text{cm}$
$\text{S} \times \text{A} - \text{AlCl}_3$		15963 $\mu\text{S}/\text{cm}$
$\text{S} \times \text{A} - \text{NaCl}$		7360 $\mu\text{S}/\text{cm}$
B - $\text{HC}_2\text{H}_3\text{O}_2$		595 $\mu\text{S}/\text{cm}$
B - $\text{HCl}$		23876 $\mu\text{S}/\text{cm}$
B - $\text{H}_3\text{PO}_4$		9058 $\mu\text{S}/\text{cm}$
$\text{S} \times \text{B} - \text{H}_3\text{BO}_3$		203 $\mu\text{S}/\text{cm}$
C - $\text{H}_2\text{O}_{\text{distilled}}$	X	203 $\mu\text{S}/\text{cm}$
C - $\text{H}_2\text{O}_{\text{tap}}$	X	562 $\mu\text{S}/\text{cm}$
C - $\text{CH}_3\text{OH}$ (Methanol)		203 $\mu\text{S}/\text{cm}$
C - $\text{C}_2\text{H}_6\text{O}_2$ (Ethylene glycol)		203 $\mu\text{S}/\text{cm}$

## DATA ANALYSIS / QUESTIONS

- How many moles of water are there in one liter (1000 mL). Assume 1 mL = 1 gram. What is the molarity of water? Is this consistent with what you wrote above?
- Based on your conductivity values, do the Group A compounds appear to be molecular, ionic, or molecular acids? Would you expect them to partially dissociate, completely dissociate, or not dissociate at all?
- Why do the Group A compounds, each with the same concentration (0.05 M), have such large differences in conductivity values? Write an equation for the dissociation of each. Explain.



# Lab Conductivity Quizz With Answers

**LM Reder**





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