

# Properties of Acids and Bases

(Adapted from Flinn Scientific Acid - Base Test Kit I - #AP4567)

## Introduction

Battery acid, stomach acid, acid rain - just a few acids in our everyday life! What does it mean when something is acidic? How do acids differ from bases? In this lab, the physical and chemical properties of a variety of acids and bases will be examined.

## Concepts

- Acids vs. Bases
- Indicators
- pH

## Background

Acids and bases make up two groups of substances that can be categorized by their physical and chemical properties. Let's take a look at the distinguishing properties:

1. **Taste:** Acids, such as lemons or oranges, taste sour. Bases, such as soap, taste bitter. *Note:* Taste should never be used to identify a lab chemical or unknown substance. This lab does not include a test using taste.
2. **Feel:** Acid solutions do not feel much different than water; however, they sting if they contact broken skin. Base solutions have a slippery feeling. *Note:* Lab chemicals or unknown substances should never be touched with the bare skin. This lab does not include a test using feel.
3. **Indicators:** An indicator is a chemical compound, either on a test paper or in a solution, that changes color depending on the acidity or basicity of a solution and, thus, is used to test for the presence of acids or bases. There are many different kinds of indicators. Litmus paper and phenolphthalein indicator solution are two of the most common.
  - Blue litmus paper turns red when dipped in an acid solution of pH 1-6.
  - Red litmus paper turns blue when dipped in a basic solution of pH 8-14.
  - In a neutral solution with pH 7, each paper retains its original color.
  - Phenolphthalein indicator solution turns a pink color in a basic solution but remains colorless (sometimes cloudy) in an acidic or neutral solution.
4. **Reaction with carbonates:** While most bases do not react with carbonates, acids react with carbonates to release carbon dioxide gas. For example:  
$$\text{BaCO}_3 + 2\text{HCl} \rightarrow \text{BaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$$
5. **Reaction with metals:** While most bases do not react with metals, acids react with active metals to release hydrogen gas. For example:  
$$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$$

# Lab 20 Properties Of Acids And Bases

**Charles Robert Sanger**



## Lab 20 Properties Of Acids And Bases:

**Chemical Laboratory Manual** Lotta Jean Bogert,1927      **National Note-book Sheets for Laboratory Work in Chemistry** Arthur Stone Dewing,1909      **Laboratory Manual for Principles of General Chemistry** J. A. Beran,Mark Lassiter,2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don t write experiments A student s scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition s experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment s Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section      *Green and Sustainable Processing* Ponnadurai Ramasami,2021-10-25 Based on The Virtual Conference on Chemistry and its Applications VCCA 2020 Research and Innovations in Chemical Sciences Paving the Way Forward held in August 2020 and organized by the Computational Chemistry Group of the University of Mauritius The chapters reflect a wide range of fundamental and applied research in the chemical sciences and interdisciplinary subjects      *Laboratory* ,1867

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web the distinction will become clear as we proceed with the study the die designer originates designs of dies used to stamp and form parts from sheet metal assemble parts together and perform a variety of other operations in this introduction you will learn basic meanings and the names of various die components then operations that are

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