

## FUNDAMENTAL IDEAS

## Atoms and Atomic Structure

All substances are made of atoms. Atoms contain three sub-atomic particles: protons, neutrons and electrons.

	Relative Charge	Relative Mass
Proton	+1 (Positive charge)	1
Neutrons	0 (No charge)	1
Electrons	-1 (Negative charge)	1/2000 (Negligible)

At the centre of an atom is its nucleus. The nucleus contains protons and neutrons. Protons are positively charged, and neutrons have no charge, so the overall charge on an atomic nucleus is positive. Electrons surround the nucleus. Electrons are negatively charged. The number of protons in an atom is equal to the number of electrons in it, so the charges cancel each other out, so there is no overall charge on an atom.

The atomic number of an element is the number of protons in its atom. The mass number of an element is the sum of the neutrons and protons in its atom. Electrons are not counted in the mass number because their relative mass is very small.

## Chemical Bonding

Most substances are made up of many different elements. Different atoms chemically bond together to form compounds. Chemical bonding involves giving, taking or sharing electrons, to form ions or molecules. The atoms react to attain the structure of a noble gas, a complete outermost shell.

Non-metals and metals react by transferring electrons. This is called ionic bonding. The metal atom will give electrons to the non-metal atom. So the metal will become a positively charged ion. The non-metal will become a negatively-charged ion. The oppositely-charged ions will be attracted to each other.

Non-metals react with each other by sharing electrons. This is called covalent bonding. The outermost shells overlap, so they share electrons, forming covalent bonds. Molecules are formed.

## Elements and the Periodic Table

A substance made up of only one type of atom is called an element. There are about a hundred different elements found naturally on Earth. Elements combine in different ways to form different substances. An atom of each element is represented by a chemical symbol, for example, O represents oxygen. The periodic table is a table of all the known elements, sorted in order of increasing atomic number. The columns in the periodic table are called groups. All the elements in a group have the same number of electrons in the highest energy level (outermost shell), and have similar properties.

## Chemical Equations

Chemical reactions can be represented by word equations or by symbol equations. Equations tell us the reactants and the products in a chemical reaction. Symbol equations are used so anyone can read them, regardless of language, and also as they show the number of each type of atom in the reaction.

In a chemical reaction, atoms simply get swapped around. No atoms are lost or made during a chemical reaction so the mass of the products equals the mass of the reactants. You can check if a chemical equation is balanced by counting the number of each type of atom on each side of the equation. If they are equal, then the equation is balanced.

## Electronic Structure

Electrons surrounding the nucleus in an atom are arranged in shells, or energy levels. The first and lowest energy level of an atom can hold only 2 electrons. The second energy level holds up to 8 electrons. The third energy level holds up to 8 electrons, and so on. The way electrons are arranged in atoms can be represented by writing an electronic structure, or drawing an electron shell diagram. You need to know the electronic structures and shell diagrams for the first 20 elements. Examples:

Hydrogen	1	
Neon	2,8	
Calcium	2,8,8,2	

## Balancing Chemical Equations

! Higher Tier !

A common exam question may ask you to balance an equation. You need to write the equation so that there are the same numbers of each type of atom on either side of the equation. But you cannot change the formula for the compounds themselves - you can only change the number of each compound. For example, hydrogen and oxygen form water. The unbalanced equation is:  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ . In this case, the number of oxygen atoms is unbalanced. The easiest way to balance it would be to make the product  $\text{H}_2\text{O}_2$ , but that would be changing the compound itself, which you cannot do. There are two oxygen atoms on the left, but only one on the right. The only thing we can do is have two water molecules. So now we have  $\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . Now the number of oxygen atoms is balanced, but there are two hydrogens on the left, and four on the right. We can fix this by having two hydrogen molecules. We now have:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . There are the same numbers of hydrogen and oxygen atoms on each side. The equation is now balanced. Try balancing this:  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ .

# Gcse Aqa Chemistry C1 Core Revision Guide

**Christine Horbury**

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