

A Microscale Titration

Preparation: Microscale techniques involving the titration of a solution of a weak acid with a solution of a strong base.

Microcalculations:

1. Prepare a series of solutions of the weak acid:

100 mL 0.100 M

100 mL 0.050 M

100 mL 0.025 M

2. Titrate each of the solutions with a strong base:

100 mL 0.100 M

3. Prepare a series of solutions of the strong base:

100 mL 0.100 M

100 mL 0.050 M

100 mL 0.025 M

4. Calculate the pH of each solution and the average pH of the solutions.

average pH = 8.8

5. Calculate the pH of the solutions of the strong base and the average pH of the solutions.

average pH = 8.8

6. Calculate the pH of the solutions of the weak acid and the average pH of the solutions.

average pH = 8.8

Microscale Titration Lab Answers

Frank Eshelman



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