

20 Balanced Chemical Equations

1. $2\text{Na (s)} + 2\text{H}_2\text{O (l)} \rightarrow 2\text{NaOH (aq)} + \text{H}_2 \text{ (g)}$
2. $\text{Zn(s)} + \text{H}_2\text{SO}_4 \text{ (aq)} \rightarrow \text{ZnSO}_4 \text{ (aq)} + \text{H}_2 \text{ (g)}$
3. $\text{CaC}_2 \text{ (s)} + 2\text{H}_2\text{O (l)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \text{Ca(OH)}_2 \text{ (aq)}$
4. $2\text{NO (g)} + \text{O}_2 \text{ (g)} \rightarrow 2\text{NO}_2 \text{ (g)}$
5. $\text{Mg (s)} + 2\text{H}_2\text{O (l)} \rightarrow \text{Mg(OH)}_2 \text{ (aq)} + \text{H}_2 \text{ (g)}$
6. $3\text{Fe (s)} + 4 \text{H}_2\text{O (g)} \rightarrow \text{Fe}_3\text{O}_4 \text{ (g)} + 4\text{H}_2 \text{ (g)}$
7. $\text{CH}_4 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{CO (g)} + 3\text{H}_2 \text{ (g)}$
8. $\text{CuO (s)} + \text{H}_2 \text{ (g)} \rightarrow \text{Cu (s)} + \text{H}_2\text{O (l)}$
9. $3\text{H}_2 \text{ (g)} + \text{N}_2 \text{ (g)} \rightarrow 2\text{NH}_3 \text{ (g)}$
10. $\text{Pb (s)} + 4\text{Na (s)} + 4\text{C}_2\text{H}_5\text{Cl (g)} \rightarrow 4\text{NaCl (s)} + \text{Pb(C}_2\text{H}_5)_4 \text{ (l)}$
11. $\text{TiCl}_4 \text{ (l)} + 4\text{Na (s)} \rightarrow 4\text{NaCl (s)} + \text{Ti (s)}$
12. $2\text{NaOH (aq)} + \text{CO}_2 \text{ (g)} \rightarrow \text{Na}_2\text{CO}_3 \text{ (aq)} + \text{H}_2\text{O (l)}$
13. $\text{Cl}_2 \text{ (g)} + 6\text{NaOH (aq)} \rightarrow \text{NaClO}_3 \text{ (aq)} + 5\text{NaCl (aq)} + 3\text{H}_2\text{O (l)}$ (for concentrated aqueous NaOH)
14. $\text{Na}_2\text{SiO}_3 \text{ (s)} + 5\text{SiO}_2 \text{ (s)} \rightarrow \text{Na}_6\text{Si}_6\text{O}_{18} \text{ (s)} + 5\text{O}_2 \text{ (g)}$
15. $\text{Na}_2\text{SiO}_3 \text{ (s)} + 4\text{C (s)} \rightarrow \text{Na}_2\text{S (s)} + 4\text{CO (g)}$
16. $\text{NaCl (aq)} + \text{NH}_4\text{OH (aq)} + \text{CO}_2 \text{ (g)} \rightarrow \text{NaHCO}_3 \text{ (s)} + \text{NH}_4\text{Cl (aq)}$
17. $2\text{NaHCO}_3 \text{ (s)} \rightarrow \text{Na}_2\text{CO}_3 \text{ (s)} + \text{H}_2\text{O} + \text{CO}_2 \text{ (g)}$
18. $\text{Na}_2\text{CO}_3 \text{ (s)} + 10\text{H}_2\text{O} \rightarrow \text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O (s)}$
19. $\text{Na}_2\text{CO}_3 \text{ (aq)} + \text{Ca(OH)}_2 \text{ (aq)} \rightarrow 2\text{NaOH (aq)} + \text{CaCO}_3 \text{ (s)}$
20. $\text{Fe}_3\text{O}_4 \text{ (s)} + 3\text{C (s)} \rightarrow 2\text{Fe (s)} + 3\text{CO (g)}$
21. $2\text{Cu(NO}_3)_2 \text{ (s)} \rightarrow 2\text{CuO (s)} + 4\text{NO}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$
22. $2\text{Al (s)} + 6\text{HCl (aq)} \rightarrow 2\text{AlCl}_3 \text{ (aq)} + 3\text{H}_2 \text{ (g)}$
23. $2\text{Al (s)} + 2\text{NaOH (aq)} + 6\text{H}_2\text{O (l)} \rightarrow 2\text{NaAl(OH)}_4 \text{ (aq)} + 3\text{H}_2 \text{ (g)}$

Lab 20 Balanced Chemical Equations

Tobias Bleicker



Lab 20 Balanced Chemical Equations:

Unveiling the Magic of Words: A Report on "**Lab 20 Balanced Chemical Equations**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Lab 20 Balanced Chemical Equations**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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