

SCIENCE (52) CHEMISTRY

SCIENCE Paper - 2

Aims:

1. To acquire the knowledge of terms, concepts, processes, techniques and principles related to the subject.
2. To develop the ability to apply the knowledge of contents and principles of chemistry in unfamiliar situations.
3. To acquire skills in proper handling of apparatus and chemicals.
4. To develop scientific temper, attitude and problem solving skills.
5. To recognize Chemical Science as having an important impact on the environment relating to cycles in nature; natural resources, pollution.

CLASS IX

There will be one paper of two hours duration of 80 marks and Internal Assessment of practical work carrying 20 marks.

The paper will be divided into two sections, Section I (40 marks) and Section II (40 marks).

Section I (compulsory) will contain short answer questions on the entire syllabus.

Section II will contain six questions. Candidates will be required to answer any four of these six questions.

Note: All chemical reactions should be studied with reference to the reactants, products, conditions, observations and the (balanced) equation.

1. Matter and its Composition: Law of Conservation of mass

- (i) Explanation of change of state of matter on the basis of Kinetic Theory of Matter.

Main postulates of Kinetic Theory of Matter and explanation of change of state on the basis of Inter-particle space and Inter-particle attraction and collision.

- (ii) Law of Conservation of Mass

Statement and explanation with examples.

2. Study of Gas Laws

- (i) The behaviour of gases under changes of temperature and pressure; explanation in

terms of molecular motion (particles, atoms, molecules); Boyle's Law and Charles' Law; absolute zero; gas equation; simple relevant calculations.

The behaviour of gases under changes of temperature and pressure; explanation in terms of molecular motion (particles, atoms, molecules). Boyle's Law (statement, mathematical form, simple calculations).

Charles' Law; (statement, mathematical form, simple calculations). Absolute zero; Kelvin scale of temperature. Gas equation $P_1 V_1 / T_1 = P_2 V_2 / T_2$; simple relevant calculations based on gas equation.

- (ii) Relationship between Kelvin Scale and Celsius Scale of temperature; Standard temperature and pressure.

Conversion of temperature from Celsius Scale to Kelvin scale and vice versa. Standard temperature and pressure. (simple calculations).

3. Elements, Compounds and Mixtures

- (i) General characteristics and differences between elements, compounds and mixtures.

Reasons for considering a substance as an element, compound or mixture may be given to make the concepts clear.

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Mike Jess



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