

Three organic compounds contain carbon, hydrogen and oxygen only.

(a) An organic compound B has the following composition by mass:

$$\text{C, 66.7\%}; \text{H, 11.1\%}; \text{O, 22.2\%}$$

Calculate the empirical formula of compound B.

$$\text{C: } 66.7 \rightarrow \frac{66.7}{12} = 5.558 \rightarrow 5.56 \rightarrow 11$$

$$\text{H: } 11.1 \rightarrow \frac{11.1}{1} = 11.1 \rightarrow 11$$

$$\text{O: } 22.2 \rightarrow \frac{22.2}{16} = 1.387 \rightarrow 1.39 \rightarrow 1$$

$$\text{C: } 11 \rightarrow 1.16 \rightarrow 1$$

$$\text{H: } 11 \rightarrow 1.16 \rightarrow 1$$

$$\text{O: } 1 \rightarrow 1.16 \rightarrow 1$$



(b) Compound B has the empirical formula $\text{C}_2\text{H}_2\text{O}$ and a relative molecular mass of 44.

Calculate the molecular formula of compound B.

$$\text{M}_{\text{emp}} = 44 \rightarrow 44$$

$$\text{M}_{\text{rel}} = 44$$

$$\text{O} = \frac{44}{44} = 1$$

$$\frac{\text{relative mass}}{\text{O} = 16} = \frac{44}{16} = 2.75 \rightarrow 3 \quad \text{molecular formula} = \text{C}_6\text{H}_6\text{O}_3$$

$$\left[\begin{array}{l} 3 \times \text{C}_2\text{H}_2\text{O} = \text{C}_6\text{H}_6\text{O}_3 \\ \text{empirical formula} \end{array} \right.$$

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Michael Rajnik



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