Question 1

1(6)

1(a) Balance (top loading), spatula, stirring rod, dropper, pipette (10 cm³), pipette (25 cm³), glass (filter) funnel, burette.

Burette readings (cm³)	Titration 1	Titration 2	Titration 3
Final volume	2.00	1.85	1.50
Initial volume	21.95	21.85	19.50
Volume of solution used	19.95*	20.00*	18.00

- 1(c)(i) See table above.
- 1(c)(ii) Average volume = (volume used in titration 1 + volume used in titration 2)/2. Average volume = ((19.95 + 20.00)/2) cm³. i.e. 19.98 cm³. Actual accurate value to be used in subsequent calculations is 19.975cm³.
- 1(d) 1000 cm³ K₂Cr₂O_{$\tau_{\rm (eq)}$} contains 0.020 mol. K₂Cr₂O_{$\tau_{\rm c}$}. 19.98 cm³ K₂Cr₂O_{$\tau_{\rm (eq)}$} contains (19.98 × (0.020/1000)) mol. K₂Cr₂O_{τ} i.e. 4.00 × 10⁻⁴ mol K₂Cr₂O_{τ}.

Actual accurate value to be used in subsequent calculations is 3.995×10^4 mol $K_2Cr_2O_2$.

- 1(e) From the balanced equation, the Fe²⁺:Cr₂O₂²⁻ stoichiometric ratio is 6:1. So 6 moles of iron ions (Fe²⁺) react with one mole of dichromate ions (Cr₂O₂²⁻).
- 1(f) Moles of iron in sample = (6 × moles of Cr₂O₂² used in titration) i.e. (6 × 4.00 × 10⁻⁶) mol. = 2.4 × 10⁻³ mol. iron.

Actual accurate value to be used in subsequent calculations is 2.397×10^{-5} mol.

1(g) Mass of iron in salt sample = (moles of iron in sample × molar mass of iron). Mass of iron in salt sample = (2.4 × 10⁻³ × 55.8) g iron i.e. 0.134 g iron.

Actual accurate value to be used in subsequent calculations is 0.1337526 g.

1(h) Percentage of iron in iron salt sample = (mass of iron in sample/mass of sample) × 100%. Percentage of iron in iron salt sample = (0.134/0.500) × 100% i.e. 26.80% iron.

If accurate values were used in all calculation steps leading to this one, the final answer would be 26.75%.

- Reduction of a species is the gain of an electron, or electrons, by that species.
- 1(j) A reducing agent is a species which causes the reduction of another species, and is itself oxidized when it reacts; it loses an electron, or electrons.

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