

## The Identity of an Insoluble Precipitate

Date: \_\_\_\_\_  
Course: \_\_\_\_\_  
Section: \_\_\_\_\_  
Instructor: \_\_\_\_\_

Student name: \_\_\_\_\_  
Team members: \_\_\_\_\_

### Prelaboratory assignment

- The following reactions, shown in unbalanced equations, are pertinent to this experiment. Balance the equations.
  - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{Ba}(\text{NH}_2\text{SO}_3)_2 + \text{HNO}_3$
  - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{BaSO}_4 + \text{NH}_4\text{NO}_3 + \text{HNO}_3$
  - $\text{Ba}(\text{NO}_3)_2 + \text{NH}_2\text{SO}_3\text{H} + \text{H}_2\text{O} \rightarrow \text{Ba}(\text{NH}_2)_2 + \text{H}_2\text{SO}_4 + \text{HNO}_3$
- When a solution containing 1.4 g of  $\text{Ba}(\text{NO}_3)_2$  and 2.4 g of  $\text{NH}_2\text{SO}_3\text{H}$  is boiled, a precipitate forms. This precipitate may be  $\text{Ba}(\text{NH}_2\text{SO}_3)_2$ ,  $\text{BaSO}_4$ , or  $\text{Ba}(\text{NH}_2)_2$ . You can find these products in the reactions whose equations you balanced in the first question.
  - Calculate the number of moles of each reactant.

# Identity Of An Insoluble Precipitate Prelab Answers

**Sabine Zange**



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