

Classify each triangle as *equilateral*, *isosceles*, or *scalene*.

3. $\triangle ABD$

equilateral

4. $\triangle BED$

Isosceles



For the triangle, find x and the measure of each side.

5. $\triangle LMN$ is an isosceles triangle, with $LM = LN$, $LM = 3x - 2$, $LN = 2x + 1$, and $MN = 5x - 2$.



$$3x - 2 = 2x + 1$$

$$x = 3$$

4.2 - Triangle Angle Sum (180 rule)

Find each measure.

6. $m\angle 1$

$$180 - 101 - 76 = 104$$

7. $m\angle 4$

$$180 - (65 + 70) = 45$$

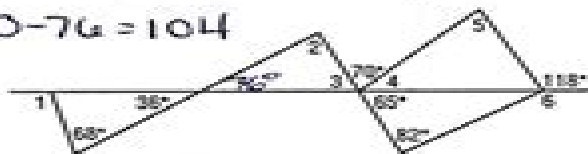
8. $m\angle 3$

$$180 - 115 = 65$$

9. $m\angle 2$

$$36 + 65 = 101$$

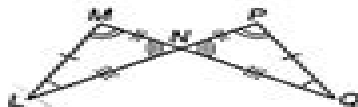
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4.3 Congruent Triangles

Show that the triangles are congruent by identifying all congruent corresponding parts. Then write a congruence statement for the triangles.

10.



$$\overline{MN} \cong \overline{PN}$$

$$\overline{NL} \cong \overline{NQ}$$

$$\overline{ML} \cong \overline{PQ}$$

$$\angle M \cong \angle P$$

$$\angle L \cong \angle Q$$

$$\angle MNL \cong \angle PNQ$$

$$\triangle MNL \cong \triangle PNQ$$

4.4 - 4.5 Congruence Shortcuts

Identify the postulate or shortcut that can be used to prove the triangles are congruent.

11.



SAS

12.



SSS

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide Ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the Ksp of a slightly soluble basic salt and how solubility changes with temperature Ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be Ksp $\text{Ca}^{2+} \text{OH}^-$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{K}_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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