

## UNIT REVIEW - BONDING

### Intramolecular Forces

1. Define electronegativity

- the tendency of atoms to attract  $e^-$

2. Electronegativity increases as you go across the periodic table, and decreases as you go down a group.

3. Explain why atoms with smaller atomic radii have stronger electronegativities

- smaller atoms means  $e^-$ 's are closer to the nucleus, which has the protons
- if you decrease the distance between protons +  $e^-$ 's, you increase attractive forces.

4. Define intramolecular force

- forces that hold atoms together: ionic, covalent, polar covalent bond

5. Compare and contrast covalent, polar covalent, and ionic bond.

Covalent	Polar covalent	Ionic
<ul style="list-style-type: none"> <li>- <math>e^-</math> are shared equally</li> <li>- between 2 non-metals</li> <li>- difference in electronegativity is <math>\leq 0.2</math></li> <li>- atoms are neutral</li> </ul>	<ul style="list-style-type: none"> <li>- <math>e^-</math> are shared unequally</li> <li>- between 2 non-metals</li> <li>- difference in electronegativity is <math>0.3 - 1.6</math></li> <li>- atoms are partially charged</li> </ul>	<ul style="list-style-type: none"> <li>- <math>e^-</math> are transferred</li> <li>- between metal + non metal</li> <li>- difference in electronegativity is <math>\geq 1.7</math></li> <li>- atoms have a charge</li> </ul>

6. What is a partial charge and how is it different from an ionic charge?

A partial charge results from an unequal sharing of electrons between 2 atoms

An ionic charge results from a transfer of electrons from 1 atom to another

# Flocabulary Unit 10 Changes Answer Key

**L Manion**



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