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Flame Test Lab

Ensential Operations: What happens when different cations are exposed to a flame?

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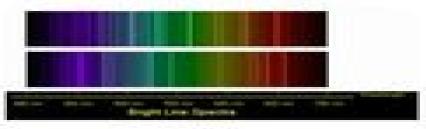
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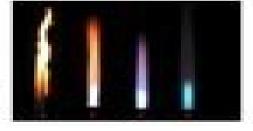
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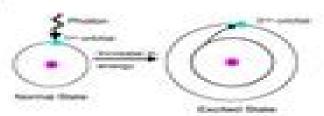


the terround intermetions

If you excite an atomor an ion by very above heating, electrons can be deprected from their normal. unescited state into higher orbitals. As they fall back down to lower levels leither in one go or in several steps), energy is released as light. Each of these tumps involves a specific amount of energy being released as light energy and each corresponds to a particular variety of the featuring. As a result of all these turnos, a spectrum of lines will be produced, some of which will be in the visible. part of the spectrum. The color you see will be a combination of all these individual colors.

In the case of sodium for other metallitors, the tumps involve very high energies and these result inlines in the UV part of the spectrum which your eyes can't see. The tamps that you can see in flame tests come translationes taling translation to a trace level

in the metal stores. So if, for example, you put sodium. chlorate which contains sodium ions, into a flame, where dothe atoms come from? In the bot flame, some of the sodium ions, edgain their electrons to form neutral sodium atoms. again. A socilium atom in an unescribed state has the structure. 1st 2st 2pt 2pt 3st, but within the flame there will be all sorts of excited states of the electrons. Bodium's tarrelar bright cognise-veltow flame optor results from promoted electrons. failing back from the Do! level to their normal Dis! level."



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