

## Section A

Answer all questions in the spaces provided

0 1

Anhydrous magnesium chloride,  $\text{MgCl}_2$ , can absorb water to form the hydrated salt  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$



← creating crystals directly is not really possible.

0 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

0 1 - 2

Some enthalpies of solution are shown in Table 1.

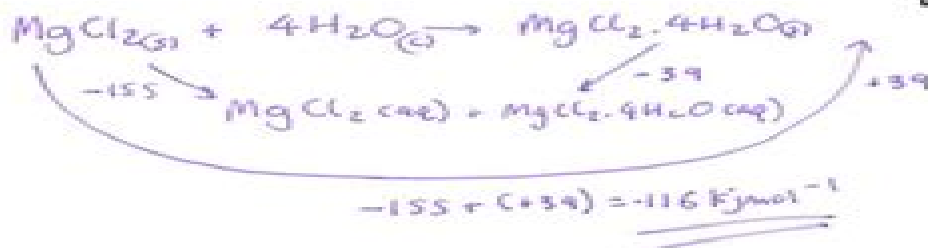
Table 1

| Salt                                                | Enthalpy of solution / $\text{kJ mol}^{-1}$ |
|-----------------------------------------------------|---------------------------------------------|
| $\text{MgCl}_2(\text{s})$                           | -155                                        |
| $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$ | -39                                         |

making the salts aqueous!

Calculate the enthalpy change for the absorption of water by  $\text{MgCl}_2(\text{s})$  to form  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$ .

[2 marks]



Enthalpy change \_\_\_\_\_  $\text{kJ mol}^{-1}$



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### **Table of Contents Max Papers Chemistry Alevel 2014**

1. Understanding the eBook Max Papers Chemistry Alevel 2014
  - The Rise of Digital Reading Max Papers Chemistry Alevel 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Max Papers Chemistry Alevel 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Max Papers Chemistry Alevel 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Max Papers Chemistry Alevel 2014
  - Personalized Recommendations
  - Max Papers Chemistry Alevel 2014 User Reviews and Ratings
  - Max Papers Chemistry Alevel 2014 and Bestseller Lists
5. Accessing Max Papers Chemistry Alevel 2014 Free and Paid eBooks

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- Max Papers Chemistry Alevel 2014 eBook Subscription Services
- Max Papers Chemistry Alevel 2014 Budget-Friendly Options
- 6. Navigating Max Papers Chemistry Alevel 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Max Papers Chemistry Alevel 2014 Compatibility with Devices
  - Max Papers Chemistry Alevel 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Max Papers Chemistry Alevel 2014
  - Highlighting and Note-Taking Max Papers Chemistry Alevel 2014
  - Interactive Elements Max Papers Chemistry Alevel 2014
- 8. Staying Engaged with Max Papers Chemistry Alevel 2014
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Max Papers Chemistry Alevel 2014
- 9. Balancing eBooks and Physical Books Max Papers Chemistry Alevel 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Max Papers Chemistry Alevel 2014
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Max Papers Chemistry Alevel 2014
  - Setting Reading Goals Max Papers Chemistry Alevel 2014
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Max Papers Chemistry Alevel 2014
  - Fact-Checking eBook Content of Max Papers Chemistry Alevel 2014
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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